

PRE-LAB QUESTIONS

1. Write names of the following monatomic ions.

Formula	Name
K^+	potassium ion
Ni^{2+}	nickel (II) ion
Zn^{2+}	zinc (II) ion
Cr^{3+}	chromium (III)
F^-	fluoride
S^{2-}	sulfide
P^{3-}	phosphide

2. Write the formulas of compounds formed by the following ions. For example, formula of compound with two
- Na^+
- ion and one
- O^{2-}
- ion is written as
- Na_2O
- . No charge should be written in the formulas of compounds.

a) compound with two Al^{3+} ions and three S^{2-} ions Al_2S_3 b) compound with one Ni^{2+} ions and two Cl^- ions $NiCl_2$

3. a) How many chloride ions (
- Cl^-
-) are needed to form a compound with a barium ion (
- Ba^{2+}
-)?
- 2

(Hint: A compound should have a net charge of zero. Therefore, sum of charges of ions in a compound should be zero.)

- b) What is the formula of compound formed containing
- Ba^{2+}
- ion and
- Cl^-
- ion?

 Cl_2Ba

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