

Practice Development

Nurse-centred evidence-based clinical education

Application for Clinical Training Form

**** indicates mandatory field****

Please Note: Applications close 8 Weeks prior to course date.

Return completed form to: Administration Officer, Nursing Midwifery Practice Development, Via Email: SC-MNPD_Training@health.qld.gov.au or
Mailing Address: Block 4, Sunshine Coast Hospital, P.O. Box 567, Nambour, Qld 4560 or Fax 07 5479 6071

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Course Details

****Name Of Course:** _____

Preferred Session In Order Of Priority

****1st Date Preference:** _____ **Time:** _____

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Participant Details

****First Name:** _____ ****Surname:** _____

****Workplace Address:** _____ ****Current Position:** _____

****Will you be attending workshop in (please circle):** **PO Leave** / **Work Time** / **Own Time**

****Contact Details: Mobile (for SMS Reminders):** _____

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****Employer Signature:** _____ ****Date:** _____

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Candidate's Name: _____ Exp. Date: _____ Signature: _____

Confirmation Of Course Attendance Is Required From Nursing Midwifery Practice Development
Confirmation will be sent via email to participant and manager.

If you have any further queries regarding course content please do not hesitate to contact Kirsten Gough on 5479 5346 or
Sharon Canny on 5479 5265.

For all course booking enquiries please contact the Administration Team, Nursing/Midwifery Practice Development on
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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should
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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the
ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $ph = -\log 0.100 = 1.000$ ph $-\log 0.100 = 1.000$ b titrant volume 12.50 ml

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web titration curves source openstax chemistry atoms first 2e strong acid strong base titrations when titrating a strong acid such as hydrochloric acid with sodium hydroxide you are reacting the HCl directly with NaOH $HCl(aq) + NaOH(aq) \rightarrow H_2O(l) + NaCl(aq)$

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web ab acid base titration curves purpose the purpose of this lab is to determine the shape of acid base titration curves theoretically and experimentally and to use an acid base titration to measure the effectiveness of antacid tablets summary a titration curve is the plot of the pH value of a solution vs the amount of titrant added

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Volume of NaOH added (mL)	pH of the solution
4	6.4
6	6.5
9	5.9
5	5.9

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