

Experiment 25

Observations and Calculations: pH: Buffers and Their Properties

A. Determination of pH by the Use of Acid-Base Indicators

Indicator	Color with 0.1 M Solution of			
	HCl	NaH_2PO_4	$\text{HC}_2\text{H}_3\text{O}_2$	ZnSO_4
Methyl violet	blue	dark violet	dark purple	purple
Thymol blue	red	yellow	orange	yellow
Methyl yellow	red	yellow	red-orange	yellow
Congo red	purple	brown	dark purple	red-orange
Bromocresol green	yellow	green	yellow	green
pH range	1-3	2-4	2-3	2-5

Circle the observation(s) for each solution that was (were) most useful in estimating the pH range.

B. Measurement of the pH of Some Typical Solutions

Record the pH and the color observed with bromocresol green for each of the 0.1 M solutions that were tested.

	NaCl^1	Na_2CO_3^2	$\text{NaC}_2\text{H}_3\text{O}_2^3$	NaHSO_4^4
pH	4.94	10.97	7.17	1.67
Color	green	blue	blue	yellow

For any two solutions having a pH less than 6 or greater than 8, write a net ionic equation to explain qualitatively why the solution has that pH.

Solution Na_2CO_3 Equation $\text{CO}_3^{2-} + \text{H}_2\text{O} \rightleftharpoons \text{HCO}_3^- + \text{OH}^-$

Solution NaHSO_4 Equation $\text{HSO}_4^- + \text{H}_2\text{O} \rightleftharpoons (\text{SO}_4)^{2-} + \text{H}_3\text{O}^+$

Explain why the color observed with bromocresol green for each of the four solutions is reasonable, given the pH.

C. Some Properties of Buffers

Buffer system selected $\text{HCO}_3^- - \text{CO}_3^{2-}$ HB is hydrogen carbonate-carbonate (name the acid)

1. pH of buffer 9.28 $[\text{H}^+] 5.24 \times 10^{-10}$ M pK_a (by Eq. 4b) _____

2. pH of diluted buffer 9.24 $[\text{H}^+] 5.75 \times 10^{-10}$ M pK_a _____

pH after addition of 5 drops NaOH 9.34

pH after addition of 5 drops HCl 9.25

Comment on your observations in Parts 1 and 2.

(continued on following page)

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Invertebrate Oxygen-binding Proteins Jean Lamy, Josette Lamy, 1981

Determination of the Size and Shape of Protein Molecules P. Alexander, R. J. Block, 2014-05-12 Analytical Methods of Protein Chemistry Volume 3 Determination of the Size and Shape of Protein Molecules provides information pertinent to the analysis and isolation of protein This book deals with the measurement of the macromolecular properties of proteins Organized into seven chapters this volume begins with an overview of the theory and practice of the electron microscope to allow an understanding of the type of object that may be examined This text then describes the methods of making protein molecules conform to such an ideal which are the techniques of specimen preparation Other chapters consider the determinations of osmotic pressures of proteins This book discusses as well the experimental basis for the theory of the diffusion process in liquids The final chapter deals with the technical problem characteristics of light scattering This book is a valuable resource for electron microscopists protein chemists biologists physicist physico chemists scientists and research workers

Chemical Aspects of the Effects of Calcium on Nerve Membrane Properties Dennis Reinhold Hafemann, 1967 Russian Journal of Physical Chemistry, 1963

Journal of Experimental Biology, 2005

American Men of Science, 1949

Water Pollution IX C. A. Brebbia, D. Prats Rico, Y. Villacampa Esteve, 2008 Featuring papers from the Ninth International Conference on Water Pollution this volume covers coastal areas and seas lakes and rivers groundwater and aquifer issues oil spills agricultural contamination environmental monitoring and sensing and remote sensing applications

Dangerous Properties of Industrial Materials Report, 1991 Journal of the American Chemical Society American Chemical Society, 1960

Proceedings of the Society are included in v 1 59 1879 1937 The Journal of Biological Chemistry, 1965 Vols 3 140 include the society s Proceedings 1907 41

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