- 5
- (b) A student carries out an experiment to measure the enthalpy change of neutralisation.
  - Define the term enthalpy change of neutralisation.

[1]

(ii) The student measures out 35.0 cm<sup>3</sup> of 2.40 moldm<sup>-3</sup> NaOH and 35.0 cm<sup>3</sup> of 2.40 mol dm<sup>-3</sup> HCI. The temperature of each solution is the same.

The student mixes the two solutions. The temperature rises by 16.5 °C.

The specific heat capacity of the mixture is 4.18 J g<sup>-1</sup> K<sup>-1</sup>. Assume that the density of the mixture is 1.00 g cm<sup>-3</sup>.

Calculate the enthalpy change of neutralisation, in kJ mol-1.

enthalpy change of neutralisation = ......kJmol<sup>-1</sup> [3]

(iii) The student repeats the experiment using 70.0 cm<sup>3</sup> of 1.20 mol dm<sup>-3</sup> HCl instead of 35.0 cm<sup>3</sup> of 2.40 mol dm<sup>-3</sup> HCl.

Explain why the temperature rise is less, and predict what the temperature rise will be.

.....

.....[2]

[Total: 11]

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