

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 \text{ kPa} / 1 \text{ atm}$ to get
 $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87°C , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

or $P = 166.29 \text{ kPa}$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56°C ,

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 329 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2916.1 \text{ L}$$

- 5) If I have an unknown quantity of gas at a pressure of 0.75 atm, a volume of 100 liters, and a temperature of 30°C , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.75 \text{ atm} \cdot 100 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 303 \text{ K}}$$

$$n = 2.91 \text{ mol}$$

- 6) If I have 2.5 moles of gas at a pressure of 0.85 atm and at a temperature of 30°C , what is the volume of the gas?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{2.5 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 303 \text{ K}}{0.85 \text{ atm}}$$

$$V = 73.1 \text{ L}$$

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Table of Contents J Weston Walch Gas Law Problems Answers

1. Understanding the eBook J Weston Walch Gas Law Problems Answers
 - The Rise of Digital Reading J Weston Walch Gas Law Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying J Weston Walch Gas Law Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an J Weston Walch Gas Law Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from J Weston Walch Gas Law Problems Answers
 - Personalized Recommendations
 - J Weston Walch Gas Law Problems Answers User Reviews and Ratings
 - J Weston Walch Gas Law Problems Answers and Bestseller Lists
5. Accessing J Weston Walch Gas Law Problems Answers Free and Paid eBooks
 - J Weston Walch Gas Law Problems Answers Public Domain eBooks
 - J Weston Walch Gas Law Problems Answers eBook Subscription Services
 - J Weston Walch Gas Law Problems Answers Budget-Friendly Options

6. Navigating J Weston Walch Gas Law Problems Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - J Weston Walch Gas Law Problems Answers Compatibility with Devices
 - J Weston Walch Gas Law Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of J Weston Walch Gas Law Problems Answers
 - Highlighting and Note-Taking J Weston Walch Gas Law Problems Answers
 - Interactive Elements J Weston Walch Gas Law Problems Answers
8. Staying Engaged with J Weston Walch Gas Law Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers J Weston Walch Gas Law Problems Answers
9. Balancing eBooks and Physical Books J Weston Walch Gas Law Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection J Weston Walch Gas Law Problems Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine J Weston Walch Gas Law Problems Answers
 - Setting Reading Goals J Weston Walch Gas Law Problems Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of J Weston Walch Gas Law Problems Answers
 - Fact-Checking eBook Content of J Weston Walch Gas Law Problems Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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