Match the expression with the logarithm that has the same value.

4. 
$$-3 \log \frac{1}{3}$$

A. 
$$\log \frac{2}{5}$$

Use log 4 - 0.602 and log 7 - 0.845 to evaluate the logarithm.

**6.** 
$$\log \frac{7}{4}$$

9. 
$$\log \frac{1}{4}$$

10. 
$$\log \frac{49}{64}$$

Expand the expression.

**12.** 
$$\log \frac{2x}{5}$$

14. 
$$\log_2 \frac{x^2}{4}$$

17. 
$$\ln \frac{1}{2x^2}$$

18. 
$$\log_9 \frac{2\pi^3}{3}$$

19. 
$$\log_6 \frac{30^2}{\sqrt{z}}$$

Condense the expression.

**20.** 
$$\log_1 4 + \log_1 2 + \log_1 2$$

**21.** 
$$\log 4 + 3 \log x + \log y$$

**22.** 
$$\log 3 + \frac{1}{2} \log x - \log 5$$

**23.** 
$$2 \ln x - \ln 3 + \ln 6$$

**24.** 
$$3 \log x + \log 4 - \log x - \log 6$$

**25.** 
$$3 \ln(x + 1) - 2 \ln y + \ln y + \ln 2$$

Use the change-of-base formula to evaluate the logarithm. Round your result to three decimal places.

In Exercises 32-34, use the following information.

Henderson-Hasselbach Formula The pH of a patient's blood can be calculated using the Henderson-Hasselbach Formula,  $pH = 6.1 + \log \frac{B}{C}$ , where B is the concentration of

bicarbonate and C is the concentration of carbonic acid. The normal pH of blood is approximately 7.4.

32. Expand the right side of the formula.

33. Find the pH of blood that has bicarbonate concentration of 38 and carbonic acid concentration of 2.0?

34. Is the pH in Exercise 33 above normal or below normal?

#### Allowbers 2

# **Holy Mcdougal Algebra 1 Cumulative Test**

**PT Brinkman** 

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