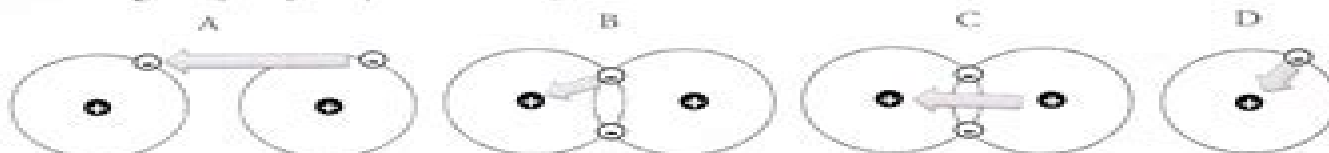


## Read This!

**Electronegativity** is a measure of the ability of an atom's nucleus to attract electrons from a different atom within a covalent bond. A higher electronegativity value correlates to a stronger pull on the electrons in a bond. This value is only theoretical. It cannot be directly measured in the lab.

12. Using the definition stated in the *Read This!* box above, select the best visual representation for electronegativity. Explain your reasoning.



13. Locate the electronegativity values in Model 1.

- a. What is the trend in electronegativity going down a group in Model 1?

*Electronegativity decreases as you go down a group (vertical column).*

- b. Explain the existence of the trend described in part a in terms of atomic structure and Coulombic attraction.

*The ability to attract electrons is based upon attractive force. The greater the distance between the nucleus and valence electrons, the less the attractive force. Less force means a lower electronegativity.*

- c. What is the trend in electronegativity going across a period in Model 1?

*Electronegativity increases from left to right across a period (horizontal row)*

- d. Explain the existence of the trend described in part c in terms of atomic structure and Coulombic attraction.

*The greater the number protons in an atom, in a particular period, the greater the attractive force between the nucleus and valence electrons. The greater attractive force means a higher electronegativity.*

14. The two diagrams below can summarize each of the three trends discussed in this activity. Write "atomic radius," "ionization energy," and "electronegativity" under the appropriate diagram.



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**The Princeton Review**



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