

**A-LEVEL CHEMISTRY**

**Practice Paper 1**

**MARK SCHEME**

Maximum marks: 105

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|   |     |       |                                                                                                                                                                                                                                                                                                                                      |   |
|---|-----|-------|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|---|
| 1 | (a) | (i)   | $[H^+][OH^-]$ <b>OR</b> $[H_3O^+][OH^-]$<br><i>Ignore (aq)<br/>Must have [ ] not ( )</i>                                                                                                                                                                                                                                             | 1 |
|   |     | (ii)  | $\sqrt{3.46 \times 10^{-14}}$ ( $= 1.86 \times 10^{-7}$ )<br><i>If no square root, CE= 0</i><br><b>pH = 6.73</b> <b>Must be 2dp</b>                                                                                                                                                                                                  | 2 |
|   |     | (iii) | $[H^+] = 10^{-11.26}$ ( $= 4.365 \times 10^{-12}$ <b>OR</b> $4.37 \times 10^{-12}$ )<br><i>Mark for working</i><br>$K_w = [4.365 \times 10^{-12} \text{ OR } 4.37 \times 10^{-12} \times 0.047] = 2.05 \times 10^{-13}$<br><i>Allow <math>2.05 \times 10^{-13} = 2.1 \times 10^{-13}</math><br/>Mark for answer<br/>Ignore units</i> | 1 |
|   | (b) | (i)   | $HCOOH \rightleftharpoons HCOO^- + H^+$<br><i>Must have <math>\rightleftharpoons</math> but ignore brackets.</i><br><b>OR</b> $HCOOH + H_2O \rightleftharpoons HCOO^- + H_3O^+$<br><i>Allow <math>HCO_2^-</math> or <math>CHOO^-</math> ie minus must be on oxygen, so penalise <math>COOH^-</math></i>                              | 1 |
|   |     | (ii)  | $K_a = \frac{[H^+][HCOO^-]}{[HCOOH]}$ <b>OR</b> $\frac{[H_3O^+][HCOO^-]}{[HCOOH]}$<br><i>Must have all brackets but allow ( )<br/>Must be <math>HCOOH</math> etc.<br/>Allow ecf in formulae from (b)(i)</i>                                                                                                                          | 1 |
|   |     | (iii) | <b>M1</b><br>$K_a = \frac{[H^+]^2}{[HCOOH]}$ ( $[H^+]^2 = 1.78 \times 10^{-4} \times 0.056 = 9.97 \times 10^{-6}$ )<br><i>Allow HA or HX etc.<br/>Allow <math>[H^+] = \sqrt{(K_a \times [HA])}</math> for M1</i>                                                                                                                     | 1 |
|   |     |       | <b>M2</b> $[H^+] = 3.16 \times 10^{-3}$<br><i>Mark for answer</i>                                                                                                                                                                                                                                                                    | 1 |
|   |     |       | <b>M3</b> <b>pH = 2.50</b> allow more than 2 dp but not fewer<br><i>Allow correct pH from their wrong <math>[H^+]</math> here only if square root shown but not taken, pH = 5.00 can score max 2 for M1 and M3</i>                                                                                                                   | 1 |

# May 200a Level Chemistry Paper

**Chandan K. Sen**



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