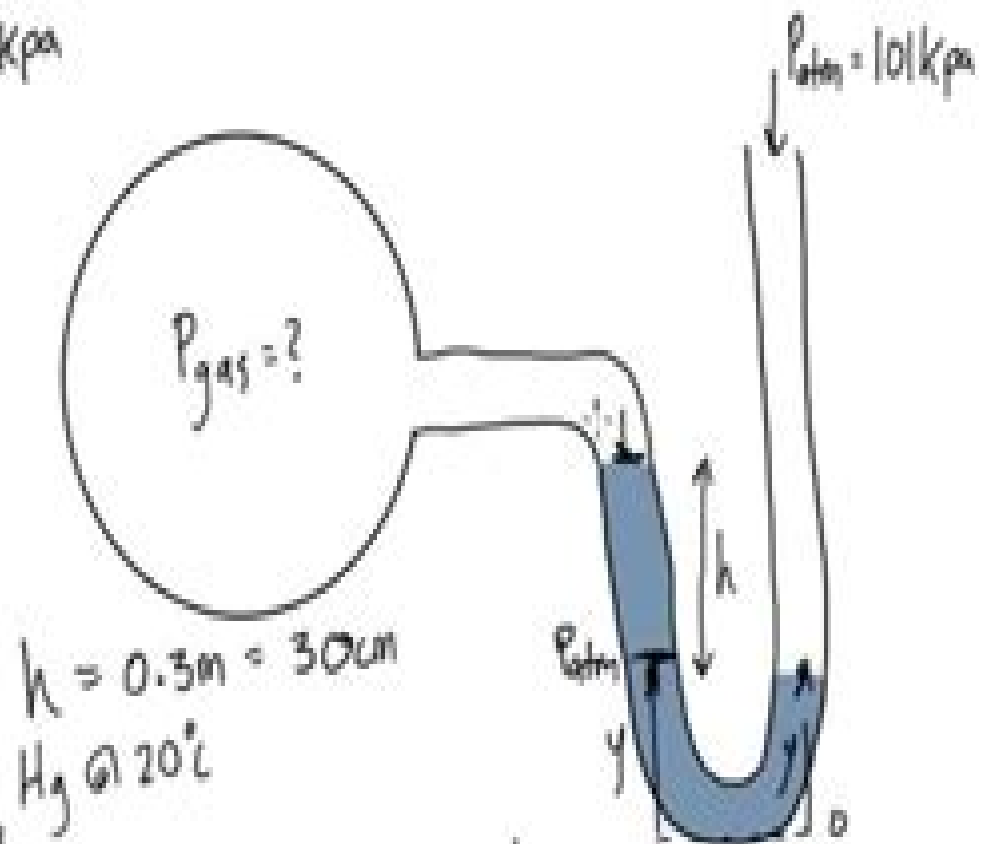


$$\begin{aligned}
 P_{gas} &= P_{atm} + \rho g h \\
 &= 101 \text{ kPa} + (13,550 \frac{\text{kg}}{\text{m}^3}) (9.81 \text{ m/s}^2) (0.3 \text{ m}) \\
 &= 101 \text{ kPa} + 39.877 \text{ kPa} \\
 P_{gas} &= 141 \text{ kPa}
 \end{aligned}$$



$h = 0.3 \text{ m} = 30 \text{ cm}$
 $\text{Hg @ } 20^\circ \text{C}$

$$\begin{aligned}
 P_{gas} &= P_{atm} - \rho g h \\
 &= 101 \text{ kPa} - (13,550) (9.81) (0.3) \\
 &= 101 - 39.877 \\
 P_{gas} &= 61 \text{ kPa}
 \end{aligned}$$

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Mark Steyvers

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