



Name:

Myriah Jackson

Date:

Student Exploration: GMOMs and the Environment

Directions: Follow the instructions below to go through the simulation. Pay special to the questions and prompts in the orange boxes.

Please do teachers and students: This lesson was designed as a follow-up to the Genetic Engineering lesson. We recommend doing that activity **before** trying this one!

Vocabularies: herbicide, insecticide, genetically modified organism, pesticide, refuge field, resistance

Prior Knowledge Questions: (Do these **BEFORE** using the Gizmo.)

- How might genetic engineering help farmers create more robust crops?

genetic engineering allows plants either to produce their own insecticides, reducing crop losses to insect damage, or to resist herbicides.

- How might GM crops affect the environment?

GM crops affect the environment in both a good and a bad way. It affects the environment by allowing more plants to increase but also affecting the surrounding areas such as the aquatic areas.

Gizmo Warm-up:

In the Genetic-Engineering Gizmo™, genetically modified (GM) corn was created that was **resistant** to insect pests and tolerant of **herbicides**. In the GMOMs and the Environment Gizmo, you will use these traits of corn to maximize corn yields while seeing how these techniques affect the environment.



On the Controls tab, make sure One-year is selected for the Mode and None is selected for the Resistance type. Make sure both sliders are set to zero. Click Play (►).

- Describe the corn field shown in the SIMULATION tab on the left.

The corn field grew at a consistent rate. When I clicked on one of the corn stalks, I saw a lot of insects eating at the corn.

- Click Resist (►). Drag the Herbicide and Insecticide sliders all the way to the right. Click Play. Describe the corn field treated with these pesticides.

Some of the corn stalks grew healthy but some of them died very early.

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Puja Mehta

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Cumulative, 1950-1977 R.R. Bowker Company. Department of Bibliography,1978 **The Publishers' Trade List Annual**,1881

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web we utilize the equation that relate amount remaining initial mass and number of half lives $n = n_0 e^{-kt}$ where n is the amount remaining, n_0 is the initial amount, k is the decay constant, and t is time. If we know the half-life $T_{1/2}$, we can express k as $k = \frac{\ln 2}{T_{1/2}}$.
125 g 5 what is the half life of an isotope that is 80 remained after 16 days solution remaining 80/100 therefore $n = 80$ $n_0 = 100$ now using the half life equation in example 1 we have

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