

CHEMISTRY I EQUATION SHEET

Solubility Rules

| Solubility Rules | | | Form a Compound which is | |
|---|---|--|--------------------------|---------|
| Negative Ion Plus Positive Ion | | | | |
| Any neg. ion + group 1a ions Soluble | | | | |
| Any neg. ion + Ammonium ion Soluble | | | | |
| Nitrate + any positive ion Soluble | | | | |
| Acetate | + | any positive ion except Ag ⁺ or Hg ²⁺ | Soluble | |
| bromide or iodide + Any other positive ion | + | Ag ⁺ , Pb ²⁺ , Hg ²⁺ or Cu ⁺ | Not Soluble | |
| Sulfate | + | Ca ²⁺ , Sr ²⁺ , Ba ²⁺ , Ra ²⁺ , Ag ⁺ or Pb ²⁺ | Not Soluble | |
| | + | Any other positive ion Soluble | | Soluble |
| Sulfide | + | group 1a ions or ammonium | Soluble | |
| | + | Be ²⁺ , Mg ²⁺ , Ca ²⁺ , Sr ²⁺ , Ba ²⁺ or Ra ²⁺ | Soluble | |
| | + | Any other positive ion | Not Soluble | |
| Hydroxide | + | group 1a ions or ammonium | Soluble | |
| | + | Any other positive ion | Not Soluble | |
| Carbonate, Chromate, Phosphate, or Sulfite + Any other positive ion | + | group 1a ions or ammonium | Soluble | |
| | + | Any other positive ion | Not Soluble | |

Equations

| | |
|--|--|
| D = m/v | c = specific heat |
| Molarity = $\frac{\text{mol}}{\text{V}}$ | d = density |
| $P_1 V_1 = P_2 V_2$ | I = initial amount |
| $\frac{T_1}{T_2}$ | F = final amount |
| Q = mc ΔT | m = mass |
| nRTPV = | M = molarity |
| F = I (1/2) * | n = number of moles & number of half lives |
| n = Total Time | P = pressure |
| Half Life Time | R = ideal gas constant |
| | Q = heat energy |
| | T = temperature |
| | V = volume |
| | X = amount of matter left |

Constants

| | |
|------------------------------------|-------------------------------------|
| | 1 atm = 101.3 kPa Pressure |
| | 1 atm = 760 mm Hg |
| | R = 8.314 L x kPa / mol°C Universal |
| Constant or 0.0821 L x atm / mol°C | |
| Avogadro's number | 6.02 X |
| Specific Heats (J / g°C) | |
| | Water = 4.180 |
| | Aluminum = 0.895 |
| | Copper = 0.387 |
| | Iron = 0.448 |

Activity Series in Decreasing Activity

| Metals |
|-----------|
| Lithium |
| Rubidium |
| Potassium |
| Calcium |
| Sodium |
| Magnesium |
| Aluminum |
| Manganese |
| Zinc |
| Iron |
| Nickel |
| Tin |
| Lead |
| Hydrogen |
| Copper |
| Silver |
| Platinum |
| Gold |
| Nonmetals |
| Fluorine |
| Chlorine |
| Bromine |
| Iodine |

Polyatomic Ions

| Name | Formula |
|-----------------------------------|---|
| Ammonium | NH ₄ ⁺ |
| Acetate | C ₂ H ₃ O ₂ ⁻ |
| Carbonate | CO ₃ ⁻² |
| Bicarbonate or Hydrogen carbonate | HCO ₃ ⁻ |
| Hypochlorite | ClO ⁻ |
| Chlorite | ClO ₂ ⁻ |
| Chlorate | ClO ₃ ⁻ |
| Perchlorate | ClO ₄ ⁻ |
| Hydroxide | OH ⁻ |
| Cyanide | CN ⁻ |
| Nitrate | NO ₃ ⁻ |
| Nitrite | NO ₂ ⁻ |
| Chromate | CrO ₄ ⁻² |
| Oxalate | C ₂ O ₄ ⁻² |
| Sulfate | SO ₄ ⁻² |
| Sulfite | SO ₃ ⁻² |
| Hydrogen sulfate | HSO ₄ ⁻ |
| Phosphate | PO ₄ ⁻³ |
| Phosphite | PO ₃ ⁻³ |
| Hydrogen phosphate | HPO ₄ ⁻² |

Metric Conversions

| Unit | Abbr. | Meaning |
|--------|-------|------------------|
| Mega- | M | 10 ⁶ |
| kilo- | k | 10 ³ |
| hecto- | h | 10 ² |
| deca- | dk | 10 ¹ |
| Liter | L | 1 |
| gram | g | 1 |
| meter | m | 1 |
| deci- | d | 10 ⁻¹ |
| centi- | c | 10 ⁻² |
| milli- | m | 10 ⁻³ |
| micro- | μ | 10 ⁻⁶ |
| nano- | n | 10 ⁻⁹ |

Calorie Conversions

| |
|--------------------|
| Water = 1 cal/ g°C |
| 1000 cal = 1 Cal |

High School Chemistry Formulas Cheat Sheet

Sebastian Brünink



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