

**AP® CHEMISTRY
2008 SCORING GUIDELINES**

Question 2

Answer the following questions relating to gravimetric analysis.

In the first of two experiments, a student is assigned the task of determining the number of moles of water in one mole of $\text{MgCl}_2 \cdot x\text{H}_2\text{O}$. The student collects the data shown in the following table.

Mass of empty container	22.347 g
Initial mass of sample and container	25.825 g
Mass of sample and container after first heating	23.983 g
Mass of sample and container after second heating	23.978 g
Mass of sample and container after third heating	23.977 g

- (a) Explain why the student can correctly conclude that the hydrate was heated a sufficient number of times in the experiment.

No additional mass was lost during the third heating, indicating that all the water of hydration had been driven off.

One point is earned for the correct explanation.

- (b) Use the data above to

- (i) calculate the total number of moles of water lost when the sample was heated, and

$$\text{mass of H}_2\text{O lost} = 25.825 - 23.977 = 1.848 \text{ g}$$

OR

$$25.825 - 23.906 = 1.849 \text{ g}$$

$$1.848 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} = 0.1026 \text{ mol H}_2\text{O}$$

One point is earned for calculating the correct number of moles of water.

- (ii) determine the formula of the hydrated compound.

$$\text{mass of anhydrous MgCl}_2 = 23.977 - 22.347 = 1.630 \text{ g}$$

$$1.630 \text{ g MgCl}_2 \times \frac{1 \text{ mol MgCl}_2}{95.20 \text{ g MgCl}_2} = 0.01712 \text{ mol MgCl}_2$$

$$\frac{0.1026 \text{ mol H}_2\text{O}}{0.01712 \text{ mol MgCl}_2} = 5.993 = 6 \text{ mol H}_2\text{O per mol MgCl}_2$$

$\Rightarrow$ Formula is $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$

One point is earned for calculating the correct number of moles of anhydrous MgCl_2 .

One point is earned for writing the correct formula (with supporting calculations).

Gravimetric Analysis Questions With Answers

Dr. (Prof.) Vitthal Vishram
Chopade, Prof. (Dr.) P.M. Patel



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Table of Contents Gravimetric Analysis Questions With Answers

1. Understanding the eBook Gravimetric Analysis Questions With Answers
 - The Rise of Digital Reading Gravimetric Analysis Questions With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gravimetric Analysis Questions With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gravimetric Analysis Questions With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gravimetric Analysis Questions With Answers
 - Personalized Recommendations
 - Gravimetric Analysis Questions With Answers User Reviews and Ratings
 - Gravimetric Analysis Questions With Answers and Bestseller Lists

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 - Highlighting and Note-Taking Gravimetric Analysis Questions With Answers
 - Interactive Elements Gravimetric Analysis Questions With Answers
8. Staying Engaged with Gravimetric Analysis Questions With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gravimetric Analysis Questions With Answers
9. Balancing eBooks and Physical Books Gravimetric Analysis Questions With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gravimetric Analysis Questions With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gravimetric Analysis Questions With Answers
 - Setting Reading Goals Gravimetric Analysis Questions With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gravimetric Analysis Questions With Answers
 - Fact-Checking eBook Content of Gravimetric Analysis Questions With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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