

Definitions

Atom -

Particles that make up all substances.

Element -

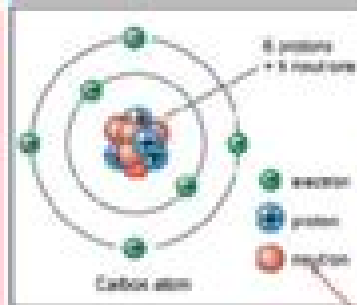
A substance made up of only one kind of atom.

Compound -

A substance made of different types of atoms joined together.

Mixture -

A substance consisting of atoms that can be easily separated as they are not joined together.



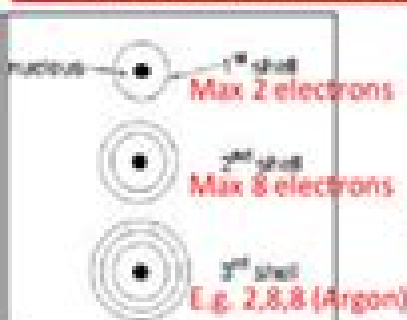
Periodic Table of the Elements																	
1	2											3	4	5	6	7	8
H	He											Li	Be	B	C	N	O
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Hf	Ta
Ra	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Mn	No	Lr	Rf	Mo
Fr	Ra	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Mn	No	Lr	Rf

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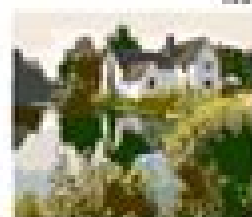
Group 0 or 8 are **UNREACTIVE**
They have a full outer shell of electrons

Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
E.g. Lithium 2,1 is a group 1 element.



Protons and neutrons both have a mass of 1



Used for building

LIMESTONE

Calcium Carbonate **CaCO₃**

HEAT
(thermal decomposition)

CaO
Calcium Oxide

Limewater

The test for CO₂

Ca(OH)₂
Or Calcium Hydroxide comes from adding water to CaO

Carbonates
Decompose when heated
= Oxide + CO₂
React with acid
= Salt + water + CO₂

Balanced Equations



C=1
H=4
O=4

A balanced equation has the same number of atoms of each element on both sides

Using Molecules: In this equation, the large number in front of the chemical symbol tells you the molecules needed.

e.g. 1 CH₄ molecule reacts with 2 O₂ molecules

Covalent Bonding

Non-metals sharing electrons



Ionic Bonding



IONS
Charged particles

TRANSITION METALS



Metals - Very useful
e.g. Copper wires conduct well
Most need to be **ALLOYED** to make them harder

Metals which are un-reactive are found in their **NATIVE STATE** e.g. **GOLD**
More reactive metals are found as **METALS ORES** and need to be **EXTRACTED**



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