

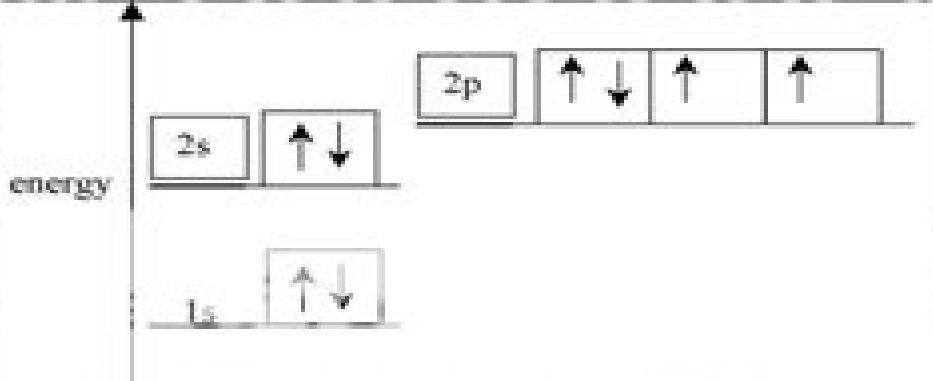


Abbreviations, annotations and conventions used in the Mark Scheme	/ = alternative and acceptable answers for the same marking point : = separates marking points NOT = answers which are not worthy of credit () = words which are not essential to gain credit <u> </u> = (underlining) key words which <u>must</u> be used to gain credit ecf = error carried forward AW = alternative wording ora = or reverse argument
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Question	Expected answers	Marks
1 (a)	  ✓ 1, 2 or 3 p orbitals are OK	[2]
(b)	d orbital 2 ✓ p sub-shell 6 ✓ 3rd shell 18 ✓	[3]
(c) (i) (ii)	 2s and 2p labels ✓ Ignore any superscripted numbers. 8 electrons in correct levels with arrows correctly shown ✓	[2]
(d) (i)	Energy change when each atom in 1 mole ✓ of gaseous atoms ✓ loses an electron ✓ (to form 1 mole of gaseous 1+ ions).	[3]
(ii)	$\text{O}^{2+}(\text{g}) \longrightarrow \text{O}^{3+}(\text{g}) + \text{e}^-$ ✓✓ 1 mark for correct species; 1 mark for state symbols No charge required on electron. Ignore (g) on e	[2]

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 type of chemical reaction for this reaction we have a neutralization reaction balancing strategies here we have a neutralization reaction

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