

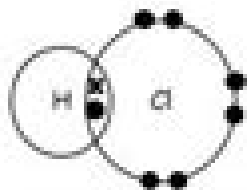
# C3 - STRUCTURE AND BONDING 3.5 Covalent bond

## Aiming for Grade 4

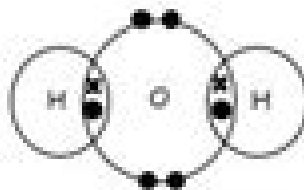
Ionic bonds form between **metals** and **non-metals**. Covalent bonds form between **non-metals**. Covalent bonds form when atoms **share** pairs of **electrons**. These bonds are very **strong**.

**Share, Metals, strong, non-metals, electrons**  
You can use the words more than once.

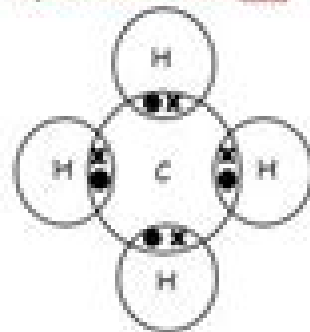
Name these small molecules and state their formulae.



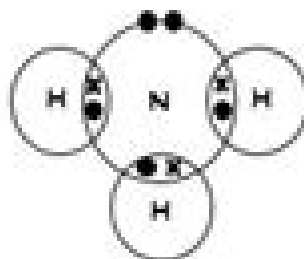
Hydrochloric Acid ( $\text{HCl}$ )



Water ( $\text{H}_2\text{O}$ )



Methane ( $\text{CH}_4$ )



Ammonia ( $\text{NH}_3$ )

Assessment

## Aiming for Grade 6

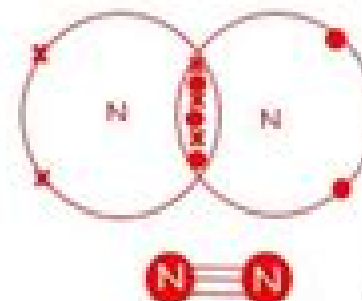
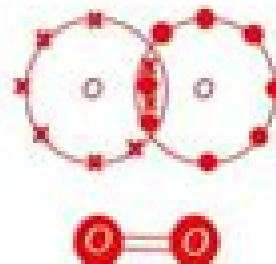
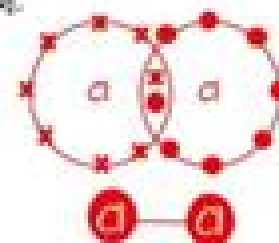
Describe how Hydrogen and Chlorine form a covalent bond to form Hydrochloric acid.

To gain full outer shells Hydrogen needs to gain an electron and Chlorine needs to lose one. Hydrogen shares its outer shell electron with Chlorine and Chlorine shares one of its electrons.

What is a double covalent bond?

Two shared pairs of electrons

Draw dot and cross diagrams and ball and stick diagrams for  $\text{H}_2$ ,  $\text{Cl}_2$ ,  $\text{O}_2$ , and  $\text{N}_2$ .



Assessment

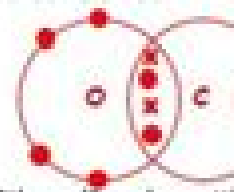
## Aiming

Draw ball and stick or the following covalent

Three hydrogen atoms  
one phosphorus (phos)



Carbon dioxide



Ethane (2 carbon with hydrogen)



How do double and triple properties?

A double or triple bond the bond making the me

Assessment  
Grade 6

1. I can draw dot and

# May 20th 2013 Chemistry C3 Mark Scheme

**Ensheng Dong**



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## **Table of Contents May 20th 2013 Chemistry C3 Mark Scheme**

1. Understanding the eBook May 20th 2013 Chemistry C3 Mark Scheme
  - The Rise of Digital Reading May 20th 2013 Chemistry C3 Mark Scheme
  - Advantages of eBooks Over Traditional Books
2. Identifying May 20th 2013 Chemistry C3 Mark Scheme
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an May 20th 2013 Chemistry C3 Mark Scheme
  - User-Friendly Interface
4. Exploring eBook Recommendations from May 20th 2013 Chemistry C3 Mark Scheme

- Personalized Recommendations
- May 20th 2013 Chemistry C3 Mark Scheme User Reviews and Ratings
- May 20th 2013 Chemistry C3 Mark Scheme and Bestseller Lists
- 5. Accessing May 20th 2013 Chemistry C3 Mark Scheme Free and Paid eBooks
  - May 20th 2013 Chemistry C3 Mark Scheme Public Domain eBooks
  - May 20th 2013 Chemistry C3 Mark Scheme eBook Subscription Services
  - May 20th 2013 Chemistry C3 Mark Scheme Budget-Friendly Options
- 6. Navigating May 20th 2013 Chemistry C3 Mark Scheme eBook Formats
  - ePub, PDF, MOBI, and More
  - May 20th 2013 Chemistry C3 Mark Scheme Compatibility with Devices
  - May 20th 2013 Chemistry C3 Mark Scheme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of May 20th 2013 Chemistry C3 Mark Scheme
  - Highlighting and Note-Taking May 20th 2013 Chemistry C3 Mark Scheme
  - Interactive Elements May 20th 2013 Chemistry C3 Mark Scheme
- 8. Staying Engaged with May 20th 2013 Chemistry C3 Mark Scheme
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers May 20th 2013 Chemistry C3 Mark Scheme
- 9. Balancing eBooks and Physical Books May 20th 2013 Chemistry C3 Mark Scheme
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection May 20th 2013 Chemistry C3 Mark Scheme
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine May 20th 2013 Chemistry C3 Mark Scheme
  - Setting Reading Goals May 20th 2013 Chemistry C3 Mark Scheme
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of May 20th 2013 Chemistry C3 Mark Scheme

- Fact-Checking eBook Content of May 20th 2013 Chemistry C3 Mark Scheme
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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