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Use the References to access important values if needed for this question.

For the reaction



$$\Delta H^\circ = -114.4 \text{ kJ and } \Delta S^\circ = -128.9 \text{ J/K}$$

The maximum amount of work that could be done when 1.54 moles of  $\text{HCl(g)}$  react at 279 K, 1 atm is  kJ.

Assume that  $\Delta H^\circ$  and  $\Delta S^\circ$  are independent of temperature.

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A) For  
 $\text{CH}_4\text{(g)}$   
 $\Delta H^\circ$   
The eq  
Assume

B) For  
 $\text{N}_2\text{(g)}$  +  
 $\Delta H^\circ$   
The eq  
Assume

C) For t  
 $2\text{HBr(g)}$   
 $\Delta H^\circ$   
The equ  
Assume

Expe



# Iln Work Answers

**Marie-Luce Bourguet, Jule M.  
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Projects Special Track 2 Sustainable Development and Immerse Learning in the Climate Emergency Special Track 3 Literacy Equity and Immersive Learning

**PC Mag**, 1993-10-26 PCMag.com is a leading authority on technology delivering Labs based independent reviews of the latest products and services Our expert industry analysis and practical solutions help you make better buying decisions and get more from technology

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**Statistics Catalog 2005** Neil Thomson,2004-09 Prev Gde Int Bus Stats Ronald M. Weiers,2004

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