

9th Grade IGCSE Chemistry Semester Exam Review

Unit 1: Atomic Theory

- Describe the structure of an atom in terms of electrons and a nucleus containing protons and neutrons and state the relative charges and approximate relative of these sub-atomic particles
- Define *proton number* and *nucleon number*.
- Use proton number and the simple structure of atoms to explain the basis of the Periodic Table, with special reference to the elements of proton number 1 to 20.
- Define *isotopes*.
- Describe the build-up of electrons in 'shells' and understand the significance of the noble gas electronic structures and of valency electrons
- Demonstrate understanding of the terms atom and molecule.

Practice Questions

1.

What do the nuclei in a hydrogen molecule contain?

- A** electrons and neutrons
- B** electrons and protons
- C** neutrons only
- D** protons only

2.

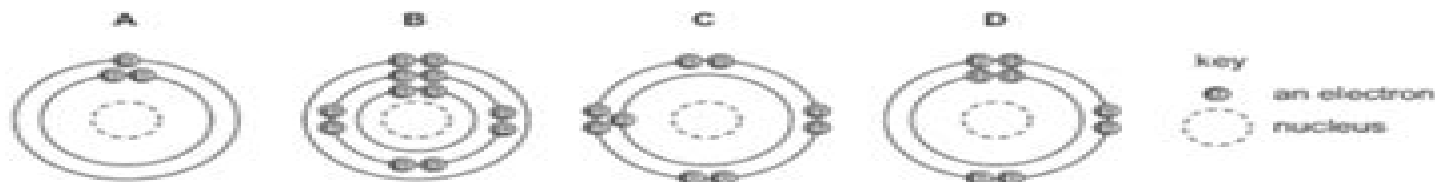
The proton number of helium is 2.

What information does this give about helium?

- A** Its atom has two electrons.
- B** Its atom is twice as heavy as a hydrogen atom.
- C** It is a Group II element.
- D** Its molecule has two atoms.

3.

Which diagram shows an atom in the same group of the Periodic Table as sodium?



4.

What is the electronic structure of an atom with a proton number 5 and a nucleon number 11?

- A** 1, 8, 2
- B** 2, 8, 1
- C** 2, 3
- D** 3, 2

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**Ian Davis, Alison Dennis, Chris
Workman**



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