

Chem 1020 Molarity worksheet, version 2

1. What is the molarity of a solution that contains 5.83 g LiCl in 1.50 L solution?

$$5.83 \text{ g LiCl} \times \frac{1 \text{ mol LiCl}}{42.389 \text{ g LiCl}} = 0.13754 \text{ mol LiCl}$$

$$M = \frac{\text{mol solute}}{\text{L solution}} = \frac{0.13754 \text{ mol LiCl}}{1.50 \text{ L solution}} = 0.0917 \text{ M}$$

2. What is the volume of a 0.300 M CuSO<sub>4</sub> solution that contains 12.4 g CuSO<sub>4</sub>?

$$12.4 \text{ g CuSO}_4 \times \frac{1 \text{ mol CuSO}_4}{159.6 \text{ g CuSO}_4} \times \frac{1 \text{ L solution}}{0.300 \text{ mol CuSO}_4} = 0.26 \text{ L solution}$$

3. What mass of solid KBr would be needed to prepare 750.0 mL of 1.75 M KBr solution? How specifically would this solution be prepared?

$$750.0 \text{ mL KBr soln} \times \frac{1 \text{ L soln}}{1000 \text{ mL soln}} \times \frac{1.75 \text{ mol KBr}}{1 \text{ L soln}} \times \frac{119.01 \text{ g KBr}}{1 \text{ mol soln}} = 156 \text{ g KBr}$$

Prepare this solution by measuring 156 g KBr into a 750 mL volumetric flask, and adding enough water to make 750 mL solution.

4. What is the % by mass of 28.3 g of solution which contains 3.56 g NaCl?

$$\% (\text{mass}) = \frac{\text{mass solute}}{\text{mass solution}} \times 100\% = \frac{3.56 \text{ g NaCl}}{28.3 \text{ g solution}} \times 100\% = 12.6\%$$

5. What volume of methanol is required to prepare 2.0 L of a 15.0% by volume solution?

$$\% (\text{vol}) = \frac{\text{vol solute}}{\text{vol solution}} \times 100\%$$

$$\text{vol solute} = \left( \frac{\% (\text{vol})}{100\%} \right) \times \text{vol solution} = \left( \frac{15.0\%}{100\%} \right) \times 2.0 \text{ L} = 0.30 \text{ L}$$

6. What volume of 12.0 M HCl is required to make 500.0 mL of 0.400 M HCl? How specifically would this solution be prepared?

$$M_1 V_1 = M_2 V_2$$

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(0.400 \text{ M})(500.0 \text{ mL})}{(12.0 \text{ M})} = 16.7 \text{ mL}$$

Prepare this solution by measuring 16.7 mL of 12.0 M HCl into a 500 mL volumetric flask, and adding water to make 500 mL of solution.

7. If 17.0 mL of 16.0 M H<sub>2</sub>SO<sub>4</sub> solution is diluted to 100.0 mL, what is the diluted concentration?

$$M_1 V_1 = M_2 V_2$$

$$M_2 = \frac{M_1 V_1}{V_2} = \frac{(16.0 \text{ M})(17.0 \text{ mL})}{(100.0 \text{ mL})} = 2.72 \text{ M}$$

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web tampa florida which has been a popular area for food trucks and hosted its first food truck event in september 2011 it also hosted a rally with 121 food trucks in march 2014 said to be a new world record breaking the prior record of 99 set in tampa in september 2013 prior to that an april 2013 food truck parade in miami with 62

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web in flnauticalfleamarket.co 954 205 7813 south florida fairgrounds west palm beach fl mark your calendars for february 10 11 2024 as under the sun promotions inc welcomes back the highly anticipated palm beach marine flea market and seafood festival to the vibrant south florida fairgrounds this incredible two day event is

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web 3 4th of july celebration downtown avalon park 2 festivals avalon park there was a decent number of food trucks and food stands to select from more 4 winter park harvest festival 1 local flavor

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