

Acids & Bases Study Guide

6. Water is considered neutral. Describe what this means in terms of $[H^+]$ & $[OH^-]$ concentrations.

8. Complete the following:

$$pH = -\log[H^+]$$

$$K_w = [H^+][OH^-]$$

7. Fill in the blank:

9. What is the $[OH^-]$ of a solution with a pH of 10?

10. What is the $[H^+]$ of a solution with a pH of 4?

11. What is the pH of a solution with a $[H^+]$ of 1×10^{-5} M?

12. What is the pH of a solution with a $[OH^-]$ of 1×10^{-4} M?

13. What is the pH of a solution with a $[H^+]$ of 1×10^{-12} M?

14. What is the pH of a solution with a $[OH^-]$ of 1×10^{-10} M?

15. What is the pH of a solution with a $[H^+]$ of 1×10^{-11} M?

16. What is the pH of a solution with a $[OH^-]$ of 1×10^{-13} M?

Name: _____ Period: _____ Date: _____

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1. Fill in the blanks or complete and contrast the properties of Acids & Bases.

| Acids | Bases |
|---|---|
| <ul style="list-style-type: none"> • Taste _____ • Feel _____ • Turn _____ paper _____ • Turn _____ paper _____ • React with _____ | <ul style="list-style-type: none"> • Taste _____ • Feel _____ • Turn _____ paper _____ • Turn _____ paper _____ • React with _____ |

2. Fill in the chart for the following acids & bases.

| Acid or Base? | Strong or Weak? | Formula | Name |
|---------------|-----------------|-----------|-------------------|
| | | | hydrochloric acid |
| | | | barium hydroxide |
| | | $HClO_4$ | |
| | | $NaOH$ | |
| | | H_2SO_4 | hydrochloric acid |

3. Predict the products for the following double replacement reactions involving acids & bases. Identify which component is the Acid, Base, Conjugate Acid (CA), and Conjugate Base (CB) in the formula.



4. Label the pH Scale from 0 (Most Acidic) to 14 (Most Basic). Identify the pH of the following solutions.



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