

Final Exam Ver. 1 (90 Multiple Choice)

Part A: Multiple Choice: (90 marks)

Identify the letter of the choice that best completes the statement or answers the question.

1. Identify the compound that is likely to have the highest boiling point?

- a. CH_4
- b. CH_3OH
- c. CH_3CH_3
- d. $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$
- e. $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$

2. What forms when an alcohol is dehydrated?

- a. an ester
- b. an acid
- c. an alkene
- d. an amine
- e. a ketone

3. What results when a secondary alcohol is oxidized?

- a. ketone
- b. alkene
- c. aldehyde
- d. acid
- e. amine

4. Both copper (atomic number 29) and chromium (atomic number 24) appear to break the pattern in the order of filling the 3d and 4s orbitals. This change in pattern is expressed by

- a. an increase in the number of electrons in both the 3d and 4s orbitals.
- b. a reduction in the number of electrons in both the 3d and 4s orbitals.
- c. a reduction in the number of electrons in the 3d orbital and an increase in the 4s orbital.
- d. a reduction in the number of electrons in the 4s orbital and an increase in the 3d orbital.

5. An electron has the following set of quantum numbers: $n = 3$, $l = 1$, $m_l = 1$, $m_s = +\frac{1}{2}$. In which orbital is this electron found?

- a. 3s
- b. 3p
- c. 3d
- d. 3f
- e. 4p

6. Which of the following groups of ions and atoms is comprised of species having exactly the same ground state electron configuration?

- a. Al^{3+} , O^{2-} , Ne, Cl
- b. Ca, Ti^{2+} , Cl⁻, S^{2-}
- c. B, He, Li, Be^{2+}
- d. Ca^{2+} , Ti^{2+} , Cl⁻, S^{2-}
- e. Ne, Ar, Kr, Xe

7. Which condensed electron configuration represents the element, Bohrium ($Z = 107$)?

- a. $[\text{Rn}][5f^{14}6f^{14}7d^2]$
- b. $[\text{Rn}][5f^{14}6d^{10}7d^3]$
- c. $[\text{Rn}][5f^{14}6f^{14}5d^3]$
- d. $[\text{Rn}][5f^{14}6f^{14}6d^3]$
- e. $[\text{Rn}][5f^{14}6f^{14}6d^2]$

8. Which type of hybridization is most likely to have occurred for the central atom in PCl_3 ?

- a. sp^3d^2
- b. sp^3d
- c. sp^3
- d. sp^2
- e. sp

9. Which electron configuration represents an alkali metal?

- a. $1s^2 2s^2 2p^6 3s^2 3p^2$
- b. $1s^2 2s^2 2p^6 3s^2 3p^5$
- c. $1s^2 2s^2 2p^6 3s^1$

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