

Name: _____
 Period: _____

Metals vs. Non-Metals; Dot Diagrams; Ions

Ch. 19.1

Metals versus Non-Metals

The dividing line is the red or bold black line from between Boron and Aluminum down and to the right. Everything to the left is a metal; to the right, non-metal. One exception (don't ya just hate that?!) is hydrogen – a non-metal.

Sodium (Na) is a metal.

Oxygen is a non-metal.

Metal or Non-metal?	
Potassium: _____	Fluorine: _____
Bromine: _____	Hydrogen: _____
Beryllium: _____	Silver: _____
Helium: _____	Nitrogen: _____

*Metals are on the left side. Non-metals on the right.
 Metals tend to lose electrons. Non-metals gain them tight.*

1 H									2 He
3 Li	4 Be		5 B	6 C	7 N	8 O	9 F	10 Ne	
11 Na	12 Mg		13 Al	14 Si	15 P	16 S	17 Cl	18 Ar	
19 K	20 Ca		31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr	

Non-metal (exception)

Dividing line

Non-Metals

Metals

Easy to remember by which side iron (Fe) is on—the left side!

On either side of the divide are the metalloids or semi-metals—they have characteristics of both B, Si, Ge, As.

Dot Diagrams

Dot Diagrams (sometimes known as Lewis dot diagrams) are a depiction of an atom's valence electrons. They are a powerful tool in helping you understand, see, and even predict molecular bonding.

Put the correct number of valence electrons around the chemical symbol in this order in pairs:



Neon has 8 valence electrons (in the right-hand column).
 Neon has no unoccupied spaces.
 It is full.

Notice that the electrons are in pairs, not a circle.

Why would the electrons spread out around the atom instead of bunching up?

Draw the Dot Diagrams for the following elements:

Beryllium

Sodium

Helium

Carbon

Boron



Lithium, with 1 valence electron (first column).

Carbon has 4 valence electrons

Oxygen has 6 valence electrons.
 How many more before it's full?

Metals Vs Nonmetals Dot Diagrams Ions Worksheet

Answer Key

R Barnett



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