

THE IDEAL GAS LAW

P = pressure in atmosphere

V = volume in liters

n = number of moles of gas

R = Universal Gas Constant = 0.0821 L.atm/mol.K

T = Kelvin temperature

See
scratch
paper

- How many moles of oxygen will occupy a volume of 2.50 liters at 1.20 atm and 25°C ?
 $PV = nRT$
 $n = \frac{PV}{RT} = \frac{1.20 \text{ atm} \cdot 2.50 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 298 \text{ K}} = 0.0122 \text{ mol}$
- What volume will 2.00 moles of nitrogen occupy at 720 mmHg and 20°C ?
 $V = \frac{nRT}{P} = \frac{2.00 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 293 \text{ K}}{0.947 \text{ atm}} = 5.00 \text{ L}$
- What pressure will be exerted by 25.0 g of CO_2 at temperature of 25°C and a volume of 500 mL?
 $n = \frac{25.0 \text{ g}}{44.01 \text{ g/mol}} = 0.568 \text{ mol}$
 $P = \frac{nRT}{V} = \frac{0.568 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 298 \text{ K}}{0.500 \text{ L}} = 2.75 \text{ atm}$
- At what temperature will 5.00 g of Cl_2 exert a pressure of 900 mmHg at a volume of 750 mL?
 $n = \frac{5.00 \text{ g}}{70.91 \text{ g/mol}} = 0.0704 \text{ mol}$
 $T = \frac{PV}{nR} = \frac{1.18 \text{ atm} \cdot 0.750 \text{ L}}{0.0704 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}} = 250 \text{ K}$
- What is the density of NH_3 at 800 mmHg and 25°C ?
 $V = 1.00 \text{ L}$
 $n = \frac{PV}{RT} = \frac{1.05 \text{ atm} \cdot 1.00 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 298 \text{ K}} = 0.0434 \text{ mol}$
 $\text{mass} = n \cdot M = 0.0434 \text{ mol} \cdot 17.03 \text{ g/mol} = 0.739 \text{ g}$
 $\text{density} = \frac{\text{mass}}{V} = 0.739 \text{ g/L}$
- If the density of a gas is 1.2 g/L at 745 mmHg and 20°C , what is its molar mass?
 $V = 1 \text{ L}$
 $n = \frac{PV}{RT} = \frac{0.980 \text{ atm} \cdot 1.00 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 293 \text{ K}} = 0.0406 \text{ mol}$
 $\text{molar mass} = \frac{\text{mass}}{n} = \frac{1.2 \text{ g}}{0.0406 \text{ mol}} = 29.5 \text{ g/mol}$
- How many moles of nitrogen gas will occupy a volume of 347 mL at 6680 mmHg and 27°C ?
 $n = \frac{PV}{RT} = \frac{8.79 \text{ atm} \cdot 0.347 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 300 \text{ K}} = 0.122 \text{ mol}$
- What volume will 454 grams (1 lb) of hydrogen occupy at 1.05 atm and 25°C ?
 $n = \frac{454 \text{ g}}{2.016 \text{ g/mol}} = 225 \text{ mol}$
 $V = \frac{nRT}{P} = \frac{225 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 298 \text{ K}}{1.05 \text{ atm}} = 5250 \text{ L}$
- Find the number of grams of CO_2 that exert a pressure of 785 mmHg at a volume of 32.5 L and a temperature of 32°C .
 $n = \frac{PV}{RT} = \frac{1.03 \text{ atm} \cdot 32.5 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 305 \text{ K}} = 1.32 \text{ mol}$
 $\text{mass} = n \cdot M = 1.32 \text{ mol} \cdot 44.01 \text{ g/mol} = 58.1 \text{ g}$
- An elemental gas has a mass of 10.3 g. If the volume is 58.4 L and the pressure is 758 mmHg at a temperature of 25°C , what is the gas?
 $n = \frac{PV}{RT} = \frac{0.997 \text{ atm} \cdot 58.4 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 298 \text{ K}} = 2.45 \text{ mol}$
 $\text{molar mass} = \frac{\text{mass}}{n} = \frac{10.3 \text{ g}}{2.45 \text{ mol}} = 4.20 \text{ g/mol}$
 The gas is H_2 .

Ideal Gas Law Chemistry If8766 Answer Key

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