

electrochemistry - the study of electricity - related applications of oxidation - reduction reactions

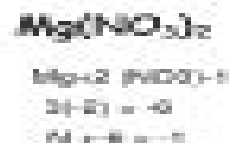
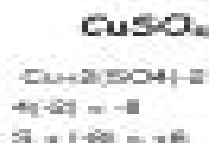
oxidation numbers -

- numbers assigned to the elements in a chemical reaction that give the apparent charge of the elements
- In ionic compounds - oxidation # = true charge
- In molecular compounds - oxidation # is based on electron distribution
- Oxidation #s are written with the sign first at and then the number

Rules For Assigning Oxidation Numbers:
situation

		oxidation #	
<u>uncombined element</u>		0	
<u>sum in a compound</u>		0	
<u>sum in a polyatomic ion</u>		charge on ion	
<u>elements in a compound:</u>			
Fluorine		-1	
Oxygen		-2	
Hydrogen		+1	
group 1		+1	
group 2		+2	
Aluminum		+3	

Example:



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Tao Wei



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