

## Naming Acids

What makes an acid unique and how are acids named?

### Why?

A variety of acids are used in foods, industry, and research. Acids are covalently bonded molecules, but when they are put into water they produce ions. One of the ions produced is always  $H^+$ , which immediately combines with a water molecule to form the hydronium ion ( $H_3O^+$ ). The  $H_3O^+$  ion is what defines the acidic properties of a substance. Because of their special classification, acids have a naming system different from ionic or other molecular (covalent) compounds.

### Model 1 – Binary Acids

Acid	Name of acid in aqueous solution	Cation (+)*	Anion (–)
HCl	Hydrochloric acid	$H_3O^+$	$Cl^-$
HBr	Hydrobromic acid	$H_3O^+$	$Br^-$
$H_2S$	Hydrosulfuric acid	$2H_3O^+$	$S^{2-}$
HF	Hydrofluoric acid	$H_3O^+$	$F^-$

\*Hydrogen ions ( $H^+$ ) join with water molecules in solution to form hydronium ions,  $H_3O^+$ .

1. Complete the cation and anion columns of the table in Model 1. Be careful to show the charges on the ions.

2. Why does hydrosulfuric acid contain two hydrogens?

*Two hydrogens with a +1 charge are required to balance the 2– charge of the sulfide anion.*

3. Look at the formulas and names of the binary acids in Model 1.

- a. What prefix is used at the beginning of the name for all binary acids?

*Binary acids begin with the prefix “hydro-.”*

- b. What suffix is used at the end of the name for all binary acids?

*The names of binary acids end with “-ic,” followed by the word “acid.”*

4. The prefix “bi” means “two.” Propose a reason that the acids in Model 1 are all referred to as “binary” acids.

*All binary acids have two different elements or atoms in them.*



5. Write a rule for naming binary acids.

*All binary acids begin with the prefix hydro-, which precedes the name of any nonmetal other than hydrogen, in the compound. The name of this nonmetal is modified to end with the suffix -ic.*

*Example: hydro\_\_\_\_\_ic acid.*



### Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (–)	Polyatomic Anion Name
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# Naming Acids Pogil

**Richard Samuel Moog**



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