

1 Consider the isoelectronic species, Na^+ , Mg^{2+} , F^- , and O^{2-} . The correct order of increasing length of their radii is

1. $\text{F}^- < \text{O}^{2-} < \text{Mg}^{2+} < \text{Na}^+$
2. $\text{Mg}^{2+} < \text{Na}^+ < \text{F}^- < \text{O}^{2-}$
3. $\text{O}^{2-} < \text{F}^- < \text{Na}^+ < \text{Mg}^{2+}$
4. $\text{O}^{2-} < \text{F}^- < \text{Mg}^{2+} < \text{Na}^+$

2 Among the following elements, the one that is not an actinoid is:

1. Curium ($Z=96$)
2. Californium ($Z=98$)
3. Uranium ($Z=92$)
4. Terbium ($Z=65$)

3 The order of screening effect of electrons of s, p, d and f orbitals of a given shell of an atom on electrons in its outer shell is :

1. $s > p > d > f$
2. $f > d > p > s$
3. $p < d < f < s$
4. $f > p > s > d$

4 The first ionisation enthalpies of Na, Mg, Al, and Si are in the order of-

1. $\text{Na} < \text{Al} < \text{Mg} < \text{Si}$
2. $\text{Na} > \text{Mg} > \text{Al} > \text{Si}$
3. $\text{Na} < \text{Mg} < \text{Al} < \text{Si}$
4. $\text{Na} > \text{Mg} > \text{Al} < \text{Si}$

5 Among halogens, the correct order of amount of energy released in electron gain (electron gain enthalpy) is-

1. $\text{F} > \text{Cl} > \text{Br} > \text{I}$
2. $\text{F} < \text{Cl} < \text{Br} < \text{I}$
3. $\text{F} < \text{Cl} > \text{Br} > \text{I}$
4. $\text{F} < \text{Cl} < \text{Br} < \text{I}$

6 The period number in the long form of the periodic table is equal to :

1. Magnetic quantum number of any element of the period.
2. Atomic number of any element of the period.
3. Maximum principal quantum number of any element of the period.
4. Maximum azimuthal quantum number of any element of the period.

7 The correct order of size of the given species is:

1. $\text{I} > \text{I}^- > \text{I}^+$
2. $\text{I}^+ > \text{I}^- > \text{I}$
3. $\text{I} > \text{I}^+ > \text{I}^-$
4. $\text{I}^- > \text{I} > \text{I}^+$

8 Electronic configuration of four elements A, B, C and D are given below :

- A. $1s^2 2s^2 2p^3$
 B. $1s^2 2s^2 2p^4$
 C. $1s^2 2s^2 2p^6 3s^1$
 D. $1s^2 2s^2 2p^5$

The correct order of increasing tendency to gain electron is :

1. $\text{A} < \text{C} < \text{B} < \text{D}$
2. $\text{A} < \text{B} < \text{C} < \text{D}$
3. $\text{D} < \text{B} < \text{C} < \text{A}$
4. $\text{D} < \text{A} < \text{B} < \text{C}$

9 Match the correct ionization enthalpies and electron gain enthalpies of the following elements.

	Elements		ΔH_1	ΔH_2	$\Delta_{\text{eg}} H$
(i)	Most reactive non-metal	A.	419	3051	-48
(ii)	Most reactive metal	B.	1681	3374	-328
(iii)	Least reactive element	C.	738	1451	-40
(iv)	Metal forming binary halide	D.	2372	5251	+48

Codes

	A	B	C	D
1.	ii	i	iv	iii
2.	i	ii	iii	iv
3.	i	iv	iii	ii
4.	iv	i	iii	ii

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