

salts. Al ($E^\circ = -1.66$ V), Mg ($E^\circ = -2.37$ V), Fe ($E^\circ = -0.44$ V), Zn ($E^\circ = -0.76$ V).

- (a) Al > Zn > Mg > Fe (b) Al < Zn < Mg < Fe
(c) Mg > Al > Zn > Fe (d) Zn > Mg > Fe > Al

20. The stronger oxidising agent has :

- (a) greater value of standard reduction potential
(b) lesser value of standard reduction potential
(c) greater tendency to be oxidised
(d) lesser tendency to be reduced

21. Which cell represents the measurement of standard electrode potential of zinc electrode ?

- (a) $\text{Pt}(s) | \text{H}_2(0.1 \text{ bar}) | \text{H}^+(1 \text{ M}) || \text{Zn}^{2+}(1 \text{ M}) | \text{Zn}$
(b) $\text{Zn}^{2+}(1 \text{ M}) | \text{Zn}(s) || \text{H}^+(1 \text{ M}) | \text{H}_2(0.1 \text{ bar}) | \text{Pt}(s)$
(c) $\text{Pt}(s) | \text{H}_2(1 \text{ bar}) | \text{H}^+(1 \text{ M}) || \text{Zn}^{2+}(1 \text{ M}) | \text{Zn}$
(d) $\text{Zn}(s) | \text{Zn}^{2+}(1 \text{ M}) || \text{H}^+(1 \text{ M}) | \text{H}_2(1 \text{ bar}) | \text{Pt}$

22. Using the given data, the strongest reducing agent and the strongest oxidising agent respectively are :

$$E^\circ_{\text{O}_2/\text{H}_2\text{O}} = 1.33 \text{ V}, \quad E^\circ_{\text{Cl}_2/\text{Cl}^-} = 1.36 \text{ V}$$

$$E^\circ_{\text{MnO}_2/\text{Mn}^{2+}} = 1.50 \text{ V}, \quad E^\circ_{\text{Zn}^{2+}/\text{Zn}} = -0.76 \text{ V}$$

- (a) Cl^- , Mn^{2+} (b) Zn, MnO_2
(c) Mn^{2+} , Cl^- (d) Cl_2 , Cl^-

23. In the electrolysis of concentrated aqueous sodium chloride solution, the half cell reaction taking place at anode will be :



$$E^\circ_{\text{O}_2/\text{H}_2\text{O}} = 1.21 \text{ V}$$



24. When lead storage battery gets discharged :

- (a) PbSO_4 is consumed (b) H_2SO_4 is consumed
(c) Lead is formed (d) SO_2 is evolved

25. When a copper wire is dipped into silver nitrate solution, silver is precipitated because : (ISC 1998)

- (a) both silver and copper have positive reduction potentials
(b) reduction potential of silver is higher than that of copper
(c) reduction potential of copper is higher than that of silver
(d) both silver and copper have negative reduction potentials

26. The number of Faradays which will be required to reduce 4 gram equivalents of Cu^{2+} to Cu metal is :

(ISC 2005)

- (a) 1 (b) 2
(c) 3 (d) 4

27. The standard electrode potentials of four metals A, B, C, D are -3.05, -1.66, -0.40 and 0.80 volts respectively. The highest chemical activity will be shown by : (ISC 2006)

- (a) A (b) B
(c) C (d) D

28. The number of Faradays required to reduce one mole of Cu^{2+} to metallic copper is : (ISC 2009)

- (a) one (b) two
(c) three (d) four

29. The unit of equivalent conductance is : (ISC 2010)

- (a) $\text{ohm}^{-1} \text{ cm}^2 \text{ eq}^{-1}$ (b) $\text{ohm}^{-1} \text{ cm}^2 \text{ g}^{-1}$
(c) $\text{ohm cm}^2 \text{ eq}^{-1}$ (d) $\text{ohm}^{-1} \text{ mol}^{-1}$

30. When zinc granule is dipped into copper sulphate, copper is precipitated because : (ISC 2008, 11)

- (a) both copper and zinc have a positive reduction potential
(b) reduction potential of copper is higher than that of zinc
(c) reduction potential of zinc is higher than that of copper
(d) both zinc and copper have a negative reduction potential

31. The quantity of electricity required to deposit 1.15 g of sodium from molten NaCl (Na=23, Cl=35.5) is :

(ISC 2012)

- (a) 1 F (b) 0.5 F
(c) 0.05 F (d) 1.5 F

32. A current liberates 0.50 g of hydrogen in 2 hours. The weight of copper (at. wt. = 63.5) is deposited at the same time by the same current through copper sulphate solution is : (ISC 2014)

- (a) 63.5 g (b) 31.8 g
(c) 15.9 g (d) 15.5 g

33. The standard electrode potentials of four non-metals A, B, C, D are +0.54, +1.06, +1.36 and 2.87 V respectively. Most reactive and most powerful oxidising agent is :

- (a) A, B (b) C, D
(c) B, D (d) D

34. Electrochemical equivalent is the amount of substance which gets deposited from its solution on passing electrical charge equal to :

- (a) 96500 coulombs (b) 1 coulomb
(c) 60 coulombs (d) 965 coulombs

35. At which electrode reduction and at which electrode oxidation occurs :

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