

Section A

For each question there are four possible answers, A, B, C and D. Choose the one you consider to be correct.

- 1 Which equation correctly represents the enthalpy change of atomisation of iodine?

- A $\frac{1}{2} \text{I}_{2(\text{s})} \longrightarrow \text{I}_{(\text{g})}$
 B $\text{I}_{2(\text{s})} \longrightarrow 2\text{I}_{(\text{g})}$
 C $\text{I}_{2(\text{s})} \longrightarrow 2\text{I}_{(\text{s})}$
 D $\frac{1}{2} \text{I}_{2(\text{s})} \longrightarrow \text{I}_{(\text{s})}$

- 2 What is the second electron affinity of oxygen?
 Use the data below to determine your answer.

	$\Delta H \text{ kJ/mol}$
Enthalpy of formation of calcium oxide	-636
Lattice enthalpy of calcium oxide	-3401
Second ionisation energy of calcium	+1100
First ionisation energy of calcium	+590
Enthalpy of atomisation of calcium	+177
Enthalpy of atomisation of oxygen	+249
First electron affinity of oxygen	-141

- A +790 kJ/mol
 B -790 kJ/mol
 C +1 039 kJ/mol
 D -1 039 kJ/mol

- 3 Why does the enthalpy change of neutralisation of oleic acid, $\text{CH}_3(\text{CH}_2)_7\text{CH}=\text{CH}(\text{CH}_2)_7\text{CO}_2\text{H}$, with aqueous ammonia differ from the theoretical value of $-57.2 \text{ kJ/mol}^{-1}$?

- A Oleic acid and ammonia do not react in the ratio 1:1.
 B Oleic acid is unsaturated therefore its enthalpy of neutralisation is more exothermic.
 C Energy is required to ionise oleic acid and ammonia.
 D Ammonia is completely hydrolysed.

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