

KEY

Radioactive Half-Life Problems

1.) Calculate the number of half lives it would take for 100 grams of Radium-223 to decay into 3.125 grams of Radon-219.

$$100 \xrightarrow[11 \text{ days}]{①} 50 \xrightarrow[11 \text{ days}]{②} 25 \xrightarrow[11 \text{ days}]{③} 12.5 \xrightarrow[11 \text{ days}]{④} 6.25 \xrightarrow[11 \text{ days}]{⑤} 3.125$$

55

2.) How long would it take for 100 grams of Radium-223 to decay into 3.125 grams of Radon-219 if the half life of Radium-223 is 11 days?

$$100g \longrightarrow 3.125g \text{ TAKES } 5 \text{ HALF-LIVES}$$

$$5 \times 11 \text{ days} = 55 \text{ DAYS}$$

3.) Astatine-217 has a half life of 20 milliseconds. If you start with a 30 gram sample of Astatine-217 how much will be left after 100 milliseconds? Hint: how many half lives in 100 milliseconds.

$$100 \text{ ms} = 5 \text{ HALF-LIVES}$$

$$20 \overline{) 100} = 5$$

$$30g \xrightarrow{①} 15g \xrightarrow{②} 7.5g \xrightarrow{③} 3.75g \xrightarrow{④} 1.875g \xrightarrow{⑤} .9375g$$

4.) Carbon-14 has a half life of 5730 years. An archeologist discovers a piece of cloth in an old cave. If it could be determined that 3.5 half lives had passed since the cloth was made how old would the cloth be?

$$5730 \text{ YRS} \times 3.5 \text{ H.L.} = 20,055 \text{ years}$$

5.) If you start with a sample of 100 atoms and determine the half life to be 10 days would the half life be any different if you then repeated your measurement using 10,000 atoms? Explain

NO - IT WOULD JUST TAKE LONGER FOR THE SUBSTANCE TO COMPLETELY DECAY.

6.) Thorium-232 decays into Radium-228 with a half-life of 1.9 years. If you start with a 60gram sample of Thorium-232 how much would you have left after 5.7 years?

$$1.9 \overline{) 5.7} = 3 \text{ HALF-LIVES}$$

$$60g \xrightarrow{①} 30g \xrightarrow{②} 15g \xrightarrow{③} 7.5g$$

7.) Lead-212 decays into Bismuth-212 with a half-life of 11 hours. If you start with 200 atoms of Lead-212 how many atoms of Bismuth-212 would you have after 22 hours.

$$11 \overline{) 22} = 2 \text{ HALF-LIVES}$$

$$\text{lead atoms} = 200 \xrightarrow{①} 100 \xrightarrow{②} 50 = \text{lead atoms}, 50 = \text{BISMUTH-212}$$

Half Life Problems Answer Key

June Eilers, Regina Fink



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Table of Contents Half Life Problems Answer Key

1. Understanding the eBook Half Life Problems Answer Key
 - The Rise of Digital Reading Half Life Problems Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Half Life Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Half Life Problems Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Half Life Problems Answer Key
 - Personalized Recommendations
 - Half Life Problems Answer Key User Reviews and Ratings

- Half Life Problems Answer Key and Bestseller Lists
- 5. Accessing Half Life Problems Answer Key Free and Paid eBooks
 - Half Life Problems Answer Key Public Domain eBooks
 - Half Life Problems Answer Key eBook Subscription Services
 - Half Life Problems Answer Key Budget-Friendly Options
- 6. Navigating Half Life Problems Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Half Life Problems Answer Key Compatibility with Devices
 - Half Life Problems Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Half Life Problems Answer Key
 - Highlighting and Note-Taking Half Life Problems Answer Key
 - Interactive Elements Half Life Problems Answer Key
- 8. Staying Engaged with Half Life Problems Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Half Life Problems Answer Key
- 9. Balancing eBooks and Physical Books Half Life Problems Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Half Life Problems Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Half Life Problems Answer Key
 - Setting Reading Goals Half Life Problems Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Half Life Problems Answer Key
 - Fact-Checking eBook Content of Half Life Problems Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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