

Previous Years AIPMT/NEET & AIIMS Ouestions

- Pressure-temperature relationship for an ideal gas undergoing adiabatic change is $(\gamma = C_c/C_c)$ 1.
 - p_T = constant

LAIPMT 92, HARYANA PMT 01, RPET 20011

- (2) $p_{T^{-1-r}} = constant$
- (3) ptr-ttr = constant
- (4) $p_{\uparrow \uparrow \uparrow \uparrow} = constant$
- An inventor claims that when temperatures of source and sink are 127° C and 27° C respectively 2. then efficiency of engine is 26%, then :-[AIPMIT 2001]
 - It is impossible.

(2) It in possible but possibility is less

(3) It's possibility is high

- (4) Data is insufficient
- If the system takes 100 cal, heat, and releases 80 cal to sink, if source temperature is 127°C 3. find the sink temp. :-[AIIMS 2001]
 - (1) 47° C.

(2) 127° C

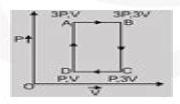
(3) 67° C

- (4) 107° C
- 4. The efficiency of carnot engine is 50% and temperature of sink is 500K. If temperature of source is kept constant and its efficiency raised to 60%, then the required temperature of the [MANIPAL 2001, AIPMT 2002] sink will be:-
 - (1) 100 K

(2) 600 K

(3) 400 K.

- (4) 500 K.
- An ideal gas heat engine operater in carnot cycle between 227°C and 127°C. It absorbs 6 × 10st 5. cals of heat at higher temperature Amount of heat converted to work is. [AIPMIT 2005]
 - (1) 2.4×10^4 cals
- (2) 6×10^4 cals
- (3) 1.2 × 104 cals
- (4) 4.8×10^4 cals
- 6. An ideal monoatomic gas is taken round the cycle ABCDA as shown in following P - V diagram. The work done during the cycle is: [RPET 92, 2000, HT 83, AIPMT 90, 92, 2007]



(1) PW

(2) 2 PV

(3) 4 PV

- (4) Zero.
- The $\left(\frac{W}{O}\right)$ of a carnot-engine is $\frac{1}{6}$, now the temp. of sink is reduced by 62°C, then this ratio 7. becomes twice, therefore the initial temp, of the sink and source are respectively :-(2) 37°C, 99°C (1) 33°C, 67°C
 - (3) 67°C, 33°C

[AIPMIT 2007]

(4) 97 K. 37 K

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Table of Contents Me121thermodynamics Previous Year Question Paper

- 1. Understanding the eBook Me121thermodynamics Previous Year Question Paper
 - The Rise of Digital Reading Me121thermodynamics Previous Year Question Paper
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Me121thermodynamics Previous Year Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Me121thermodynamics Previous Year Question Paper
 - User-Friendly Interface

- 4. Exploring eBook Recommendations from Me121thermodynamics Previous Year Question Paper
 - Personalized Recommendations
 - Me121thermodynamics Previous Year Question Paper User Reviews and Ratings
 - Me121thermodynamics Previous Year Question Paper and Bestseller Lists
- 5. Accessing Me121thermodynamics Previous Year Question Paper Free and Paid eBooks
 - Me121thermodynamics Previous Year Ouestion Paper Public Domain eBooks
 - Me121thermodynamics Previous Year Question Paper eBook Subscription Services
 - Me121thermodynamics Previous Year Question Paper Budget-Friendly Options
- 6. Navigating Me121thermodynamics Previous Year Question Paper eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Me121thermodynamics Previous Year Question Paper Compatibility with Devices
 - Me121thermodynamics Previous Year Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Me121thermodynamics Previous Year Question Paper
 - Highlighting and Note-Taking Me121thermodynamics Previous Year Question Paper
 - Interactive Elements Me121thermodynamics Previous Year Question Paper
- 8. Staying Engaged with Me121thermodynamics Previous Year Question Paper
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Me121thermodynamics Previous Year Question Paper
- 9. Balancing eBooks and Physical Books Me121thermodynamics Previous Year Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Me121thermodynamics Previous Year Question Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Me121thermodynamics Previous Year Question Paper
 - $\circ \ \ Setting \ Reading \ Goals \ Me121 thermodynamics \ Previous \ Year \ Question \ Paper$
 - Carving Out Dedicated Reading Time

- 12. Sourcing Reliable Information of Me121thermodynamics Previous Year Question Paper
 - Fact-Checking eBook Content of Me121thermodynamics Previous Year Question Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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