

GROUP 2 ELEMENTS AND THEIR COMPOUNDS

- show their outermost electron placed on s subshell
- readily react to lose 2 electrons
 - $\text{M} \rightarrow \text{M}^{2+} + 2\text{e}^{-}$

Trends down the group:

- 1st ionisation energy decreases
 - Distance between outermost electron and nucleus increases
 - More shielding
 - Counteractive effect of increasing nuclear charge
- Atomic radius increases
- Melting point generally decreases
 - Much lower than all transition metals

Physical properties:

- Good conductors of heat and electricity
 - Sea of delocalised electrons
- Form white compounds
 - Transition metals form coloured compounds
- Oxidation number of +2
- All compounds are ionic or almost ionic
 - Be compounds show some covalent character because small highly charged cation polarises anion shell of anion
- Oxides / hydroxides are basic
 - React with acids (proton acceptors)
 - Except:
 - BeO
 - Be(OH)_2 which is amphoteric
- Density = 2.83 g/cm^3
 - Average density of transition metals is about 7 g/cm^3
- Compared to group 2 metals:
 - Harder + denser
 - Higher melting point
 - Metallic bond is stronger (higher charge on cation & more electrons delocalised)

Chemical properties:

- Strong reducing agents
- Oxidation is losing electrons
- Reduction is gaining electrons:
 - Reducing agent - donates electrons (reducing other substances)
 - Oxidising agent - accepts electrons (oxidising itself)
- E.g.
 - $\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}_2$
 - $\text{ZnO} \rightarrow \text{Zn}^{2+} + 2\text{O}^{2-}$
 - $\text{C} + 2\text{O}^{2-} \rightarrow \text{CO}_2 + 4\text{e}^{-}$

June 2013 Ocr Chemistry F321

Christian Drosten



June 2013 Ocr Chemistry F321:

June 2013 Ocr Chemistry F321 Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**June 2013 Ocr Chemistry F321**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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