

What is the half-life of a 100.0 g sample of nitrogen-16 that decays to 12.5 g of nitrogen-16 in 21.6 s?

$$\ln\left(\frac{N}{N_0}\right) = \left(\frac{-0.693}{t_{1/2}}\right) t$$

$$\ln\left(\frac{12.5\text{g}}{100.0\text{g}}\right) = \left(\frac{-0.693}{t_{1/2}}\right) (21.6\text{s})$$

$$-2.08 = \left(\frac{-0.693}{t_{1/2}}\right) (21.6\text{s})$$

$$t_{1/2} =$$

# Half Life Calculations Answers Chemistry If8766

**G Psacharopoulos**



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