Gay-Lussac's Law Problem

- A steel container of nitrous oxide at 15.0 atm is cooled from 25°C to –40°C. What is the final pressure at constant V?
- We first have to convert the temp from °C to K:

$$25^{\circ}\text{C} + 273 = 298 \text{ K}$$

 $-40^{\circ}\text{C} + 273 = 233 \text{ K}$ $P_1 \times \frac{T_2}{T_1} = P_2$

$$15.0 \text{ atm} \times \frac{298 \text{ K}}{233 \text{ K}} = 11.7 \text{ atm}$$

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Wolfgang Guggemos

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