



Emd Sd 70 Operating Manual

JL Elias



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determine based on electronegativity and symmetry if the following molecules are polar or nonpolar hcn h c n cf₄ c f₄ cocl₂
c o c l₂ nf₃ n f₃ q3

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molecule is one where the charge distribution around the molecule is not symmetric it results from having polar bonds and
also a molecular structure where the bond polarities do not cancel

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unequally non polar polar unequally polar the degree of polarity of a bond between any two atoms is determined by
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bonds including covalent ionic and hydrogen bonds and london dispersion forces introduction living things are made up of
atoms but in most cases those atoms aren t just floating around individually

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molecules and ions contain polar bonds which of these molecules and ions have dipole moments a clf₅ b clo₂ clo₂ c tecl₂ 4
tecl₄ 2 d pcl₃ e sef₄ f ph₂ ph₂ g xef₂ answer problem 6 2 3 6 2 3

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web the relative electronegativities of the bonded atoms is h c cl and so the bond moments all point toward the cl end of the
molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry
such as bf₃ trigonal planar ch₄ tetrahedral pf₅ trigonal bipyramidal and sf₆

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