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# Mole and Stoichiometry Mini-Lab

01/28/2013

Substance	Mass (g)	Volume (mL)
Aluminum	13.3 g	
Water	22.8 g	22.8 mL

$$\text{g H}_2\text{O} = \frac{22.8 \text{ mL H}_2\text{O}}{1 \text{ mL H}_2\text{O}} \times \frac{1.00 \text{ g H}_2\text{O}}{1 \text{ mL H}_2\text{O}} = 22.8 \text{ g H}_2\text{O}$$

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