

change of the surrounding water. $\Delta H_{\text{metal}} = -\Delta H_{\text{water}}$ If the water temperature increases, then the metal must have released heat. We can find the enthalpy change of the metal with the formula: $\Delta H_{\text{metal}} = -m_{\text{water}} \cdot C_{\text{water}} \cdot \Delta T_{\text{water}}$

Check-for-understanding question:

1. When 25 mL of water containing 0.85 mol HCl is mixed with 40 mL of water containing .05 mol KOH in a calorimeter (with initial temperature 23 °C), the temperature of the solution increases to 25.5 °C. How many kilojoules (kJ) of heat are released by the acid-base reaction?

$$\begin{aligned}\Delta H_{\text{water}} &= m_{\text{water}} \times C_{\text{water}} \times \Delta T_{\text{water}} \\ &= 65 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C} \times 2.5^\circ\text{C} \\ &= 679.25 \text{ J} = .67925 \text{ kJ}\end{aligned}$$

Measuring And Expressing Enthalpy Changes Answers

Ensheng Dong



Measuring And Expressing Enthalpy Changes Answers:

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In some sort of taken by displays and the ceaseless chatter of instantaneous communication, the melodic beauty and mental symphony produced by the written term frequently disappear into the back ground, eclipsed by the persistent noise and interruptions that permeate our lives. Nevertheless, situated within the pages of **Measuring And Expressing Enthalpy Changes Answers** a stunning fictional value filled with fresh emotions, lies an immersive symphony waiting to be embraced. Crafted by a masterful composer of language, that charming masterpiece conducts readers on a mental trip, well unraveling the hidden songs and profound affect resonating within each carefully constructed phrase. Within the depths with this touching evaluation, we will examine the book is key harmonies, analyze its enthralling publishing fashion, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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