

(b) T is an alcohol, $C_nH_{2n}O$. A gaseous sample of T occupied a volume of 20 cm^3 at 120°C and 100 kPa .

The sample was completely burned in 200 cm^3 of oxygen (an excess). The final volume, measured under the same conditions as the gaseous sample, was 130 cm^3 .

Under these conditions, all water produced is vapour. Removal of the water vapour from the gaseous mixture decreased the volume to 110 cm^3 .

Treating the remaining gaseous mixture with concentrated alkali, to absorb carbon dioxide, decreased the volume to 10 cm^3 .

The equation for the complete combustion of T can be represented as shown.



(i) Use the data given to calculate the value of x.

$$\begin{array}{l} \text{Ratio} \\ T : O_2 \\ 20 : 60 \\ 1 : 3 \end{array}$$

$$x = 3$$

$$x = \underline{3} \quad [1]$$

(ii) Use the data given to calculate the value of y.

$$\begin{array}{l} T : H_2O \\ 20 : 80 \\ 1 : 4 \end{array}$$

$$\begin{array}{l} y/2 = 4 \\ y = 8 \end{array}$$

$$y = \underline{8} \quad [1]$$

If you were unable to calculate values for x and y then use $x = 4$ and $y = 10$ for the remaining parts of this question. These are not the correct values.

(iii) Complete the equation for the complete combustion of the alcohol, T.



(iv) Give the skeletal formulae for two possible structures of T.

Name each alcohol.

(9701-S 2017-Paper 2/3-Q4)

Question

Answer



1 2 3

(a) P is a hydrocarbon, C_nH_m . A gaseous sample of P occupied a volume of 25 cm^3 at 37°C and 100 kPa .

The sample was completely burned in 200 cm^3 of oxygen (an excess). The final volume, measured under the same conditions as the gaseous sample (so that the water produced is liquid and its volume can be ignored), was 150 cm^3 .

Treating the remaining gaseous mixture with concentrated alkali, to absorb carbon dioxide, decreased the volume to 50 cm^3 .

The equation for the complete combustion of P can be represented as shown.



(i) Use the data given to calculate the value of x.

$$\begin{array}{l} P : CO_2 \\ 25 : 50 \\ 1 : 2 \end{array} \quad x = 2$$

$$x = \underline{2} \quad [1]$$

(ii) Use the data given to calculate the value of y.

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