

# Key

## Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant  $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$  to solve the following problems:

If pressure is needed in kPa then convert by multiplying by  $101.3 \text{ kPa} / 1 \text{ atm}$  to get  
 $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of  $87^\circ\text{C}$ , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

$$\text{or } P = 166.29 \text{ kPa}$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of  $56^\circ\text{C}$ , what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 327 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

- 5) If I have 17 moles of gas at a temperature of  $67^\circ\text{C}$ , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 390 \text{ K}}{88.89 \text{ L}}$$

$$P = 5.34 \text{ atm}$$

$$\text{or } P = 540.61 \text{ kPa}$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

# Ideal Gas Law Problems Worksheet Answers

**Dr. S. Pancharatnam**



## **Ideal Gas Law Problems Worksheet Answers:**

**A Guide to Teaching in the Active Learning Classroom** Paul Baepler, J. D. Walker, D. Christopher Brooks, Kem Saichaie, Christina I. Petersen, 2023-07-03 While Active Learning Classrooms or ALCs offer rich new environments for learning they present many new challenges to faculty because among other things they eliminate the room's central focal point and disrupt the conventional seating plan to which faculty and students have become accustomed. The importance of learning how to use these classrooms well and to capitalize on their special features is paramount. The potential they represent can be realized only when they facilitate improved learning outcomes and engage students in the learning process in a manner different from traditional classrooms and lecture halls. This book provides an introduction to ALCs briefly covering their history and then synthesizing the research on these spaces to provide faculty with empirically based practical guidance on how to use these unfamiliar spaces effectively. Among the questions this book addresses are: How can instructors mitigate the apparent lack of a central focal point in the space? What types of learning activities work well in the ALCs and take advantage of the affordances of the room? How can teachers address familiar classroom management challenges in these unfamiliar spaces? If assessment and rapid feedback are critical in active learning, how do they work in a room filled with circular tables and no central focus point? How do instructors balance group learning with the needs of the larger class? How can students be held accountable when many will necessarily have their backs facing the instructor? How can instructors evaluate the effectiveness of their teaching in these spaces? This book is intended for faculty preparing to teach in or already working in this new classroom environment, for administrators planning to create ALCs or experimenting with provisionally designed rooms, and for faculty developers helping teachers transition to using these new spaces.

**Physics Workbook For Dummies** Steven Holzner, 2007-10-05 Do you have a handle on basic physics terms and concepts but your problem solving skills could use some static friction? Physics Workbook for Dummies helps you build upon what you already know to learn how to solve the most common physics problems with confidence and ease. Physics Workbook for Dummies gets the ball rolling with a brief overview of the nuts and bolts, i.e. converting measures, counting significant figures, applying math skills to physics problems, etc. before getting into the nitty gritty. If you're already a pro on the fundamentals, you can skip this section and jump right into the practice problems. There you'll get the lowdown on how to take your problem solving skills to a whole new plane without ever feeling like you've been left spiraling down a black hole. With easy-to-follow instructions and practical tips, Physics Workbook for Dummies shows you how to unleash your inner Einstein to solve hundreds of problems in all facets of physics, such as Acceleration, distance, and time; Vectors; Force; Circular motion; Momentum and kinetic energy; Rotational kinematics and rotational dynamics; Potential and kinetic energy; Thermodynamics; Electricity and magnetism. Complete answer explanations are included for all problems so you can see where you went wrong or right. Plus you'll get the inside scoop on the ten most common mistakes people make when solving physics problems and how to avoid them. When

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University,1995     Ideal Gases Lifeliqe,2019 This lesson plan covers the ideal gas law and the different values for the ideal gas constant how to make various calculations using the ideal gas law and explains the conditions under which real gases are most or least ideal     **Gas Laws** Source Wikipedia,2013-09 Please note that the content of this book primarily consists of articles available from Wikipedia or other free sources online Pages 24 Chapters Acentric factor Amagat s law Avogadro s law Boyle s law Charles s law Combined gas law Compressibility factor Dalton s law Gay Lussac s law Graham s law Henry s law Magic number chemistry Partial pressure Psychrometric constant Redlich Kwong equation of state Van der Waals constants data page Van der Waals equation Excerpt The van der Waals equation is an equation of state for a fluid composed of particles that have a non zero volume and a pairwise attractive inter particle force such as the van der Waals force It was derived in 1873 by Johannes Diderik van der Waals who received the Nobel prize in 1910 for his work on the equation of state for gases and liquids The equation is based on a modification of the ideal gas law and approximates the behavior of real fluids taking into account the nonzero size of molecules and the attraction between them The van der Waals isotherms the model correctly predicts a mostly incompressible liquid phase but the oscillations in the phase transition zone do not fit experimental data The equation uses the following state variables the pressure of the fluid  $p$  total volume of the container containing the fluid  $V$  number of moles  $n$  and absolute temperature of the system  $T$  One form of the equation is where is the volume of the container shared between each particle not the velocity of a particle is the total number of particles and is Boltzmann s constant given by the universal gas constant  $R$  and Avogadro s constant  $N_A$  Extra parameters are introduced  $a$  is a measure for the attraction between the particles and  $b$  is the average volume excluded from  $v$  by a particle The equation can be cast into the better known form where is a measure of the attraction between the particles is the volume excluded by a mole of particles A careful distinction     **Aberrations from the Ideal Gas Laws and a Precision Method for the Determination of the Densities of Gases** Douglas Le Baron Peters Cooper,1931     Ideal Gas Law, Enthalpy, Heat Capacity, Heats of Solution and Mixing Eric H Snider,1984-01-01     Gas Constant 22 Success Secrets - 22 Most Asked Questions on Gas Constant - What You Need to Know Mark Love,2014-10-19 A new Gas constant Guide that will give you all There has never been a Gas constant Guide like this It contains 22 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Gas constant A quick look inside of some of the subjects covered Units of pressure Pressure of an ideal gas Glossary of engineering I Surface tension Contact angles Atmospheric entry Real equilibrium gas model Thermodynamic equations Notation Capacitive deionization Adsorption and desorption cycles Glossary of chemistry terms U Diffusion Einstein s mobility and Teorell formula Gas constant Individual gas constants Reaction intermediate Definition Thermal inertia Gas volumetric heat capacities Ideal gas law GIS in Environmental Contamination Soil contamination in

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