Practice Test Ch 3

Stoichiometry

- Remember, this is practice Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the tenting conditions autlined below.
- DO NOT USE A CALCULATOR, AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to stary practicing without one. You may use ONLY a periodic table.
- While you should practice working as first as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the "easy math" - common factors and rough estimation - do nor do "long division" to try to get exact values. Remember it is a MC test, use the answers
- Mark which questions you would like to "go over" when we get to school in September.
- 1. Holomor the following equation:

The balanced equation shows that 1.00 mole of Nills requires __ emoker(s) self-40s

- a. 0.57
- b. 1.25
- 1.33 (30) 1.75
- 4 3.5
- 2. Write a balanced equation for the combustion of acetaldobydo, CHI/CHO.

When properly balanced, the equation indicates that mole(x) of O₂ are required for each mole of CH₂CHO.

46. S. 2 c) 25 30

> 40. 6

- Imai CHICHO(5 moi 02)
- 3. What is the total mass of products formed when 16 grams of CH₂ is burned with enters oxygen?
 - 2 12 € CH₂ * ZO₂ → CO₄ * ZH₂O THE GCHA (IMOLCHA) (IMOLCHA) (IMOLCHA) (IMOLO)
 - No. 1946 pt. 44.8
 - 祖 62 度
 - (P) 80 g
- 4. Write a balanced equation for the combustion of propane, Collins

When balanced, the equation indicates that ___ moles of Os are required for each mole of Calla.

- m. 1.5 b. 3
- ec 3.5
- (4) 5

5. Balance the following equation with the SMALLEST WHOLE NUMBER COUTTICIENTS possible. Select the member that is the sum of the coefficients in the bulmeed equation:

2 KCIO. - 2 KCI - 3 O-

- 46 4.3 7
- 6. Culculate the mass of hydrogon formed when 27 g of alternium resets with excess hydrochloric acid according to the believed equation below.

- a. 1.5 g.
- b. 2.0 g. (c) 3.0 g
- E. 6.0 a
- 100 12.6
- How many grams of nitric acid, HNOs, can be proposed. from the reaction of 138 g of NO₂ with 54.0 g (154).

- 3NO₃ + 10₄O + 2HNO₃ + NO 135 g NO2 (100 NO) (2 ma HNO) (155 HNO) (100 HNO)
- 8. 92. b. 1008
- (d.) 126 11879 279

 - $= G(GG)_{G}$

#123.64S

27g A (1mol A) (2mol Hz) (2gHz) (1mol Hz)

- 8. Which of the following statements is true?
 - The molar mass of CaCO₅ is 100.1 g mol⁻¹.
 - III. 50 g of CaCOs contains 9 ×10²³ environ atoms.
 - III. A 200 e sample of CaCOs contains 2 moles of CaCOs
 - a. I only
 - b. II only
 - III only.
 - (d.) I and III only e. I. III. and IIII.

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The Princeton Review

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