# 1st Semester Chemistry Final Exam Review

1. Identify the following reaction type AND correct products:

Sc + 3 AgOH --->

- a. combustion
- b. synthesis
- e. decomposition
- d. single replacement
- e. double replacement
- w. Ag + ScOH
- x.  $Ag + 2 Sc(OH)_2$
- y. 3 Ag + Sc(OH)3
- z. 3 Ag + 2 Sc(OH)4
- Convert: 2.32 x 10<sup>5</sup> mg = ? kg

- I. a. & w.
  II. b. & z.
  III. b. & y.
  IV. d. & z.
  V. e. & z.
  VI. d. & y.
  - VIII. e. & y.
  - VIII. c. & z.
  - IX. d. & x.

Calculate the energy for a quantum of light with a frequency of 8.44 x 10<sup>18</sup> Hz.

Determine the oxidation number for iodine in aluminum iodate.

- 5. REPORTING WITH THE CORRECT NUMBER OF SIGNIFICANT DIGITS, how many grams are
  - in 4.011 moles of calcium chloride?
  - a) 445.14078 grams b) 445.140 grams
- c) 445.1 grams
- d) 445 grams

- e) 450 grams
- f) 400 grams
- 6 Calculate the frequency of light that has a wavelength of 210 nm.

# **Honors Chemistry Practice Test**

Christopher R. Hren, John T. Moore, Peter J. Mikulecky

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