

Stoichiometry & Percent Yield Practice

- 1) Calculate the molar mass of each compound. Round to 3 decimal places.

a) NaCl 58.442 g/mol

b) FeCl_3 162.204 g/mol

c) $\text{Mg}(\text{NO}_3)_2$ 246.055 g/mol

- 2) $2 \text{AgNO}_3 + \text{Ni} \rightarrow 2 \text{Ag} + \text{Ni}(\text{NO}_3)_2$ BOTH SIGNIFICANTLY

- a) If 6.50 moles of nickel (Ni) reacts with excess silver nitrate (AgNO_3), how many moles of silver (Ag) will be produced?

$$6.50 \text{ mol Ni} \times \frac{2 \text{ mol Ag}}{1 \text{ mol Ni}} = \boxed{13.0 \text{ mol Ag}}$$

mole to mole

- b) If 2.35 moles of silver nitrate (AgNO_3) reacts completely, how many grams of silver (Ag) will be created?

$$2.35 \text{ mol AgNO}_3 \times \frac{2 \text{ mol Ag}}{2 \text{ mol AgNO}_3} \times \frac{107.868 \text{ g Ag}}{1 \text{ mol Ag}} = \boxed{253 \text{ g Ag}}$$

mole to mole molar mass

- c) If 2.13 g of Ni react with excess silver nitrate, how many moles of $\text{Ni}(\text{NO}_3)_2$ will be produced?

$$2.13 \text{ g Ni} \times \frac{1 \text{ mol Ni}}{58.693 \text{ g Ni}} \times \frac{1 \text{ mol Ni}(\text{NO}_3)_2}{1 \text{ mol Ni}} = \boxed{0.036 \text{ mol Ni}(\text{NO}_3)_2}$$

molar mass mole to mole

- d) If 100 g of Ni is produced from the reaction above, what mass of $\text{Ni}(\text{NO}_3)_2$ will also be produced?

$$100 \text{ g Ni} \times \frac{1 \text{ mol Ni}}{58.693 \text{ g Ni}} \times \frac{1 \text{ mol Ni}(\text{NO}_3)_2}{1 \text{ mol Ni}} \times \frac{246.055 \text{ g Ni}(\text{NO}_3)_2}{1 \text{ mol Ni}(\text{NO}_3)_2} = \boxed{169 \text{ g Ni}(\text{NO}_3)_2}$$

molar mass mole to mole molar mass

Holt Stoichiometry Test Answers

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Holt Chemistry Ralph Thomas Myers,2004 Children's Books in Print R R Bowker Publishing,Bowker,1999-12

Progress in Ecological Stoichiometry Dedmer B. Van de Waal,James J. Elser, Adam C. Martiny,Robert W.

Sterner,James B. Cotner,2018 Ecological stoichiometry concerns the way that the elemental composition of organisms shapes their ecology It deals with the balance or imbalance of elemental ratios and how that affects organism growth nutrient cycling and the interactions with the biotic and abiotic worlds The elemental composition of organisms is a set of constraints through which all the Earth s biogeochemical cycles must pass All organisms consume nutrients and acquire compounds from the environment proportional to their needs Organismal elemental needs are determined in turn by the energy required to live and grow the physical and chemical constraints of their environment and their requirements for relatively large polymeric biomolecules such as RNA DNA lipids and proteins as well as for structural needs including stems bones shells etc These materials together constitute most of the biomass of living organisms Although there may be little variability in elemental ratios of many of these biomolecules changing the proportions of different biomolecules can have important effects on organismal elemental composition Consequently the variation in elemental composition both within and across organisms can be tremendous which has important implications for Earth s biogeochemical cycles It has been over a decade since the publication of Sterner and Elser s book Ecological Stoichiometry 2002 In the intervening years hundreds of papers on stoichiometric topics ranging from evolution and regulation of nutrient content in organisms to the role of stoichiometry in populations communities ecosystems and global biogeochemical dynamics have been published Here we present a collection of contributions from the broad scientific community to highlight recent insights in the field of Ecological Stoichiometry

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