

Balancing Equations Worksheet

- 1) $\text{___ C}_3\text{H}_8 + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$
- 2) $\text{___ Al} + \text{___ Fe}_3\text{N}_2 \rightarrow \text{___ AlN} + \text{___ Fe}$
- 3) $\text{___ Na} + \text{___ Cl}_2 \rightarrow \text{___ NaCl}$
- 4) $\text{___ H}_2\text{O}_2 \rightarrow \text{___ H}_2\text{O} + \text{___ O}_2$
- 5) $\text{___ C}_6\text{H}_{12}\text{O}_6 + \text{___ O}_2 \rightarrow \text{___ H}_2\text{O} + \text{___ CO}_2$
- 6) $\text{___ H}_2\text{O} + \text{___ CO}_2 \rightarrow \text{___ C}_7\text{H}_8 + \text{___ O}_2$
- 7) $\text{___ NaClO}_3 \rightarrow \text{___ NaCl} + \text{___ O}_2$
- 8) $\text{___ (NH}_4\text{)}_3\text{PO}_4 + \text{___ Pb(NO}_3\text{)}_4 \rightarrow \text{___ Pb}_3\text{(PO}_4\text{)}_4 + \text{___ NH}_4\text{NO}_3$
- 9) $\text{___ BF}_3 + \text{___ Li}_2\text{SO}_3 \rightarrow \text{___ B}_2\text{(SO}_3\text{)}_3 + \text{___ LiF}$
- 10) $\text{___ C}_7\text{H}_{17} + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$
- 11) $\text{___ CaCO}_3 + \text{___ H}_3\text{PO}_4 \rightarrow \text{___ Ca}_3\text{(PO}_4\text{)}_2 + \text{___ H}_2\text{CO}_3$
- 12) $\text{___ Ag}_2\text{S} \rightarrow \text{___ Ag} + \text{___ S}_8$
- 13) $\text{___ KBr} + \text{___ Fe(OH)}_3 \rightarrow \text{___ KOH} + \text{___ FeBr}_3$
- 14) $\text{___ KNO}_3 + \text{___ H}_2\text{CO}_3 \rightarrow \text{___ K}_2\text{CO}_3 + \text{___ HNO}_3$
- 15) $\text{___ Pb(OH)}_4 + \text{___ Cu}_2\text{O} \rightarrow \text{___ PbO}_2 + \text{___ CuOH}$
- 16) $\text{___ Cr(NO}_2\text{)}_2 + \text{___ (NH}_4\text{)}_2\text{SO}_4 \rightarrow \text{___ CrSO}_4 + \text{___ NH}_4\text{NO}_2$
- 17) $\text{___ KOH} + \text{___ Co}_3\text{(PO}_4\text{)}_2 \rightarrow \text{___ K}_3\text{PO}_4 + \text{___ Co(OH)}_2$
- 18) $\text{___ Sn(NO}_2\text{)}_4 + \text{___ Pt}_3\text{N}_4 \rightarrow \text{___ Sn}_3\text{N}_4 + \text{___ Pt(NO}_2\text{)}_4$
- 19) $\text{___ B}_2\text{Br}_6 + \text{___ HNO}_3 \rightarrow \text{___ B(NO}_3\text{)}_3 + \text{___ HBr}$
- 20) $\text{___ ZnS} + \text{___ AlP} \rightarrow \text{___ Zn}_3\text{P}_2 + \text{___ Al}_2\text{S}_3$

Grade 9 Natural Science Balancing Equations

Rachel Sandford



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