

# Year 11 C2 - Discovering Chemistry Topic 1

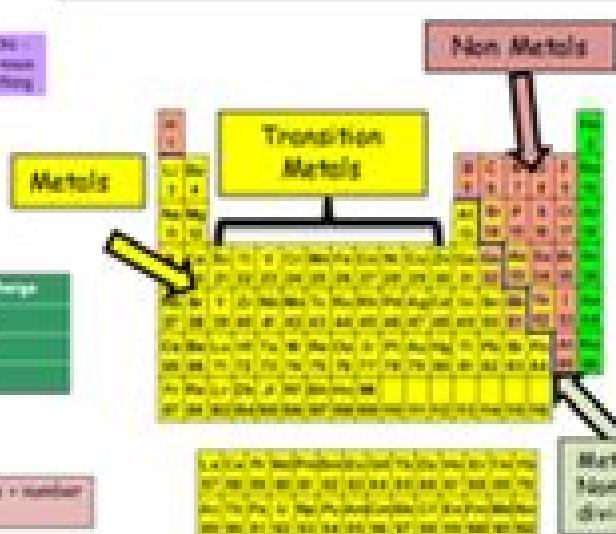


| Particle | Relative Mass | Relative Charge |
|----------|---------------|-----------------|
| Proton   | 1             | +               |
| Neutron  | 1             | 0               |
| Electron | Very small    | -               |

I came up with the first periodic table. I arranged the elements in order of increasing atomic mass

I put elements with similar properties into horizontal rows. I also left gaps incase any new elements were discovered

Dmitri Mendeleev



**14**  
**N**  
**7**

Electron structure  
+ 2, 5

**24**  
**Mg**  
**12**

Electron structure  
+ 2, 8, 2

**40**  
**Ca**  
**20**

Electron structure  
+ 2, 8, 8, 2

**He**

**4** ← **MASS NUMBER** = number of protons + number of neutrons

**2** ← **PROTON NUMBER** = number of protons (atomic)

**2** ← **CHARGE**

## Isotopes

Arrange in a table with a different number of neutrons

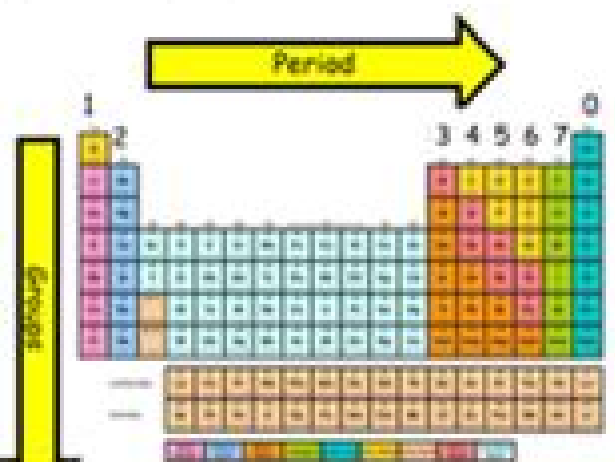
Notice that the mass number is different. How many neutrons does each isotope have?

**16**  
**O**  
**8**

**17**  
**O**  
**8**

**18**  
**O**  
**8**

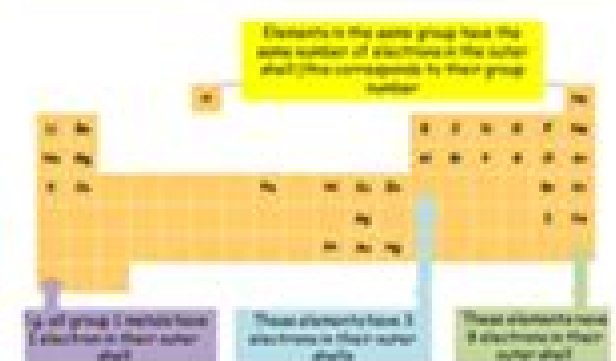
Each isotope has 8 protons - if it didn't then it just wouldn't be oxygen any more



How can an atom have a **decimal** for its mass?

This is because out of every four naturally occurring chlorine atoms, 3 have a mass of 35 and 1 has a mass of 37 so the average atomic mass is:

**$35 \times 3 + 1 \times 37 = 35.5$**



(Most important) All of the elements in the same group have similar **PROPERTIES**. This is how I thought of the periodic table in the first place. This is called **PERIODICITY**.

# Gcse Edexcel Chemistry C2 2013

**Tao Wei**



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