



Periodic Table of the Elements

Legend:

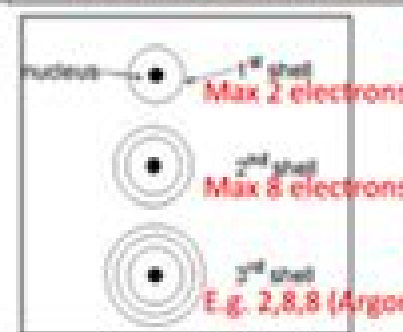
- metals
- non-metals
- metalloids
- transition metals
- alkali metals
- alkali earth metals
- transition metals
- lanthanoids
- actinoids

An atom:
Number of Protons = Number of Electrons
Mass number = Protons + Neutrons

C2 Chemistry

Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
E.g. Lithium 2,1 is a group 1 element.

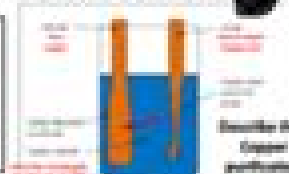


Protons and neutrons both have a mass of 1

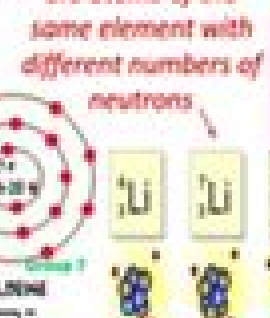
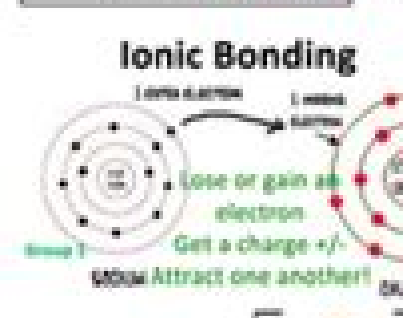
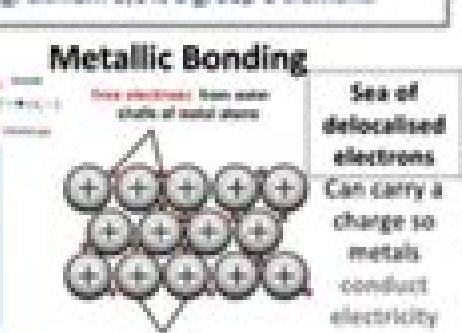
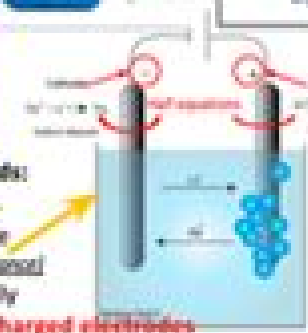
ISOTOPES
are atoms of the same element with different numbers of neutrons

OIL RIG

Oxidation is loss (of electrons)
Reduction is gain (of electrons)



Electrolysis:
Splitting up a substance using electricity



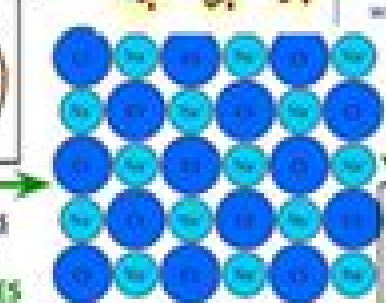
When there are 2 possible products
The less reactive forms at the electrode

Electrolysis of Brine

Used for: Disinfectants, Bleach and plastics

Used for: Margarine and HCl

Used for: Bleach, paper and soap



As a solid: High melting and Boiling Point
Conducts electricity when melted or dissolved:
Ions move freely and carry a charge

NANO-TECHNOLOGY
Means REALLY REALLY REALLY small!!!!

Metal ions only arranged in layers
ELECTROSTATIC ATTRACTION
between
+IONS and ELECTRONS

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well

Gcse Additional Chemistry C2 Revision Booklet

Neil Dixon



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