

METRIC CONVERSION CHART FOR LENGTH

Metric Length Conversions

Quantity	Units		Quantity	Units
1	centimeter (cm)	=	10	millimeters (mm)
1	meter (m)	=	100	centimeters (cm)
1	kilometer (km)	=	1000	meters (m)

Metric Length Conversions

Quantity	Units		Quantity	Units
1	kilometer (km)	=	0.04	inch (in)
1	centimeter (cm)	=	0.4	inch (in)
1	meter (m)	=	39.4	inches (in)
1	meter (m)	=	3.3	feet (ft)
1	kilometer (km)	=	1093.6	yard (yd)
1	kilometer (km)	=	0.62	miles (mi)

Metric Length Conversions

Quantity	Units		Quantity	Units
1	inch (in)	=	2.54	centimeters (cm)
1	yard (yd)	=	91.44	centimeters (cm)
1	mile (mi)	=	1609.3	meters (m)

Measurements Conversions A Complete Guide

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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton $h_a + b \rightleftharpoons h_b$ strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral pH at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $\text{pH} = -\log 0.100 = 1.000$ b titrant volume 12.50 ml

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