

## Periodic Trends Worksheet

Directions: Use the table to determine the following properties:

1. Write the following elements by increasing atomic radius: sodium, sulfur, aluminum, oxygen, potassium.

**K, Na, Al, S, O**

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, carbon, aluminum.

**Na, Al, S, C, O**

3. Which atom has the higher first ionization energy: Na or K?

**Na has a higher first ionization energy than K because it is smaller.**

4. Which two elements in the same family generally have similar properties?

**There have been two atoms that share a same valence shell, potassium and sodium.**

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

**Period 7, francium, cesium, barium, potassium, rubidium, sodium, lithium, hydrogen.**

7. Which trend in ionization energy values across periods and the periodic table? Which element has the highest?

**Increases across the periodic table, fluorine has the highest.**

8. Write the atoms in each group that have the largest atomic radius.

Li	Be
Na	Mg
K	Ca
Rb	Sr
Cs	Ba
Fr	Ra

Al	Si	P
Ga	Ge	As
In	Sn	Sb
Tl	Pb	Bi

9. Circle the atom in each group that has the greatest electronegativity.

# If8767 Answer Key Pg 45

**Annelies Wilder-Smith**



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