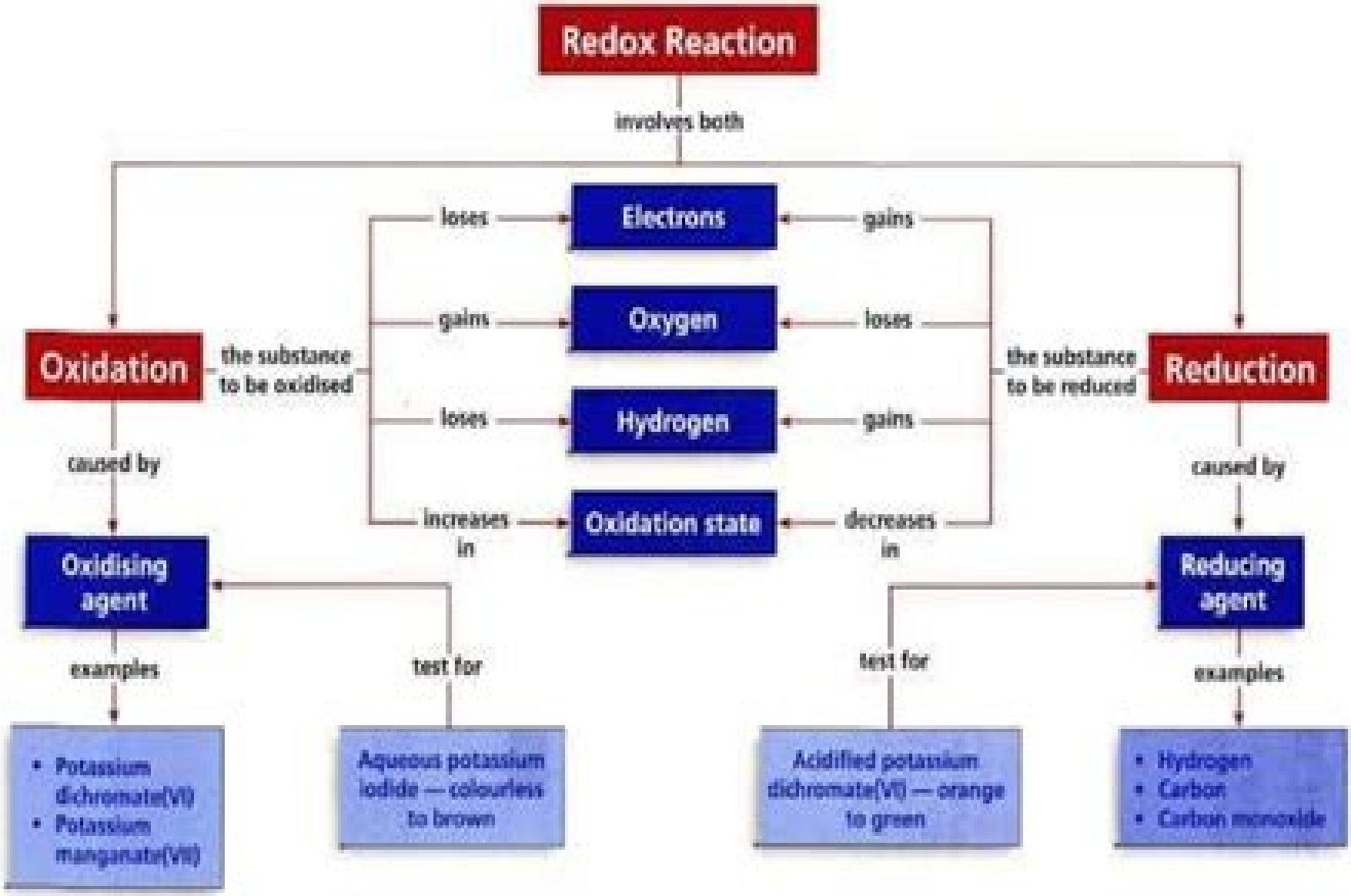
Redox Reaction
involves both
Oxidation
the substance to be oxidised
Reduction
the substance to be reduced
loses
Electrons
gains
gains
Oxygen
loses
loses
Hydrogen
gains
increases in
Oxidation state
decreases in
caused by
caused by
Oxidising agent
Reducing agent
examples
test for
test for
examples
• Potassium dichromate(VI)
• Potassium manganate(VII)
Aqueous potassium iodide — colourless to brown
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• Hydrogen
• Carbon
• Carbon monoxide

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