

The Ideal Gas Law

Introduction

In this lab report, I will be using an experiment to study the ideal gas law. I will use Boyle's Law, Charles's Law, and Avogadro's Law. I will also use the ideal gas law to find the molar mass of a gas.

The Ideal Gas Law

$$PV = nRT$$

The ideal gas law is a mathematical equation that relates the pressure, volume, and temperature of a gas to the number of particles. It is a combination of Boyle's Law, Charles's Law, and Avogadro's Law. The ideal gas law is used to find the molar mass of a gas.

The ideal gas law is used to find the molar mass of a gas. It is a combination of Boyle's Law, Charles's Law, and Avogadro's Law. The ideal gas law is used to find the molar mass of a gas.



Ideal Gas Laws Lab Escience

KJ Lindholm-Leary



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Gay Lussac's law Graham's law Henry's law Magic number chemistry Partial pressure Psychrometric constant Redlich Kwong equation of state Van der Waals constants data page Van der Waals equation Excerpt The van der Waals equation is an equation of state for a fluid composed of particles that have a non zero volume and a pairwise attractive inter particle force such as the van der Waals force It was derived in 1873 by Johannes Diderik van der Waals who received the Nobel prize in 1910 for his work on the equation of state for gases and liquids The equation is based on a modification of the ideal gas law and approximates the behavior of real fluids taking into account the nonzero size of molecules and the attraction between them The van der Waals isotherms the model correctly predicts a mostly incompressible liquid phase but the oscillations in the phase transition zone do not fit experimental data The equation uses the following state variables the pressure of the fluid p total volume of the container containing the fluid V number of moles n and absolute temperature of the system T One form of the equation is
$$\left(p + \frac{a}{V_m^2}\right)(V_m - b) = RT$$
 where V_m is the volume of the container shared between each particle not the velocity of a particle n is the total number of particles and k is Boltzmann's constant given by the universal gas constant R and Avogadro's constant N_A Extra parameters are introduced a is a measure for the attraction between the particles and b is the average volume excluded from V_m by a particle The equation can be cast into the better known form
$$\left(p + \frac{a}{V_m^2}\right)(V_m - b) = RT$$
 where a is a measure of the attraction between the particles b is the volume excluded by a mole of particles A careful distinction

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