

THE IDEAL GAS LAW

T = Kelvin temperature

See
scratch
paper

- How many moles of oxygen will occupy a volume of 2.50 liters at 1.20 atm and 25°C ?
 $PV = nRT$
 $n = \frac{PV}{RT} = \frac{1.20 \text{ atm} \cdot 2.50 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 298 \text{ K}} = 0.947 \text{ mol}$
- What volume will 2.00 moles of nitrogen occupy at 720 mmHg and 20°C ?
 $V = \frac{nRT}{P} = \frac{2.00 \text{ mol} \cdot 0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 293 \text{ K}}{0.947 \text{ atm}} = 5.00 \text{ L}$
- What pressure will be exerted by 25.0 g of CO_2 at temperature of 25°C and a volume of 500 mL?
 $n = \frac{25.0 \text{ g}}{44.0 \text{ g/mol}} = 0.568 \text{ mol}$
 $P = \frac{nRT}{V} = \frac{0.568 \text{ mol} \cdot 0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 298 \text{ K}}{0.500 \text{ L}} = 1.18 \text{ atm}$
- At what temperature will 5.00 g of Cl_2 exert a pressure of 900 mmHg at a volume of 750 mL?
 $n = \frac{5.00 \text{ g}}{71.0 \text{ g/mol}} = 0.0704 \text{ mol}$
 $T = \frac{PV}{nR} = \frac{1.18 \text{ atm} \cdot 0.750 \text{ L}}{0.0704 \text{ mol} \cdot 0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}} = 275 \text{ K}$
- What is the density of NH_3 at 800 mmHg and 25°C ?
 $V = 1.00 \text{ L}$
 $n = \frac{PV}{RT} = \frac{1.05 \text{ atm} \cdot 1.00 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 298 \text{ K}} = 0.0438 \text{ mol}$
 $\text{Density} = \frac{m}{V} = \frac{0.0438 \text{ mol} \cdot 17.0 \text{ g/mol}}{1.00 \text{ L}} = 0.745 \text{ g/L}$
- If the density of a gas is 1.2 g/L at 745 mmHg and 20°C , what is its molar mass?
 $V = 1 \text{ L}$
 $n = \frac{PV}{RT} = \frac{0.980 \text{ atm} \cdot 1.00 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 293 \text{ K}} = 0.0407 \text{ mol}$
 $M = \frac{m}{n} = \frac{1.2 \text{ g}}{0.0407 \text{ mol}} = 29.5 \text{ g/mol}$
- How many moles of nitrogen gas will occupy a volume of 347 mL at 6680 mmHg and 27°C ?
 $n = \frac{PV}{RT} = \frac{8.79 \text{ atm} \cdot 0.347 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 300 \text{ K}} = 0.125 \text{ mol}$
- What volume will 454 grams (1 lb) of hydrogen occupy at 1.05 atm and 25°C ?
 $n = \frac{454 \text{ g}}{2.02 \text{ g/mol}} = 225 \text{ mol}$
 $V = \frac{nRT}{P} = \frac{225 \text{ mol} \cdot 0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 298 \text{ K}}{1.05 \text{ atm}} = 5.00 \times 10^3 \text{ L}$
- Find the number of grams of CO_2 that exert a pressure of 285 mmHg at a volume of 32.5 L and a temperature of 32°C .
 $n = \frac{PV}{RT} = \frac{1.03 \text{ atm} \cdot 32.5 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 305 \text{ K}} = 1.32 \text{ mol}$
 $m = n \cdot M = 1.32 \text{ mol} \cdot 44.0 \text{ g/mol} = 58.1 \text{ g}$
- An elemental gas has a mass of 10.3 g. If the volume is 58.4 L and the pressure is 758 mmHg at a temperature of 25°C , what is the gas?
 $n = \frac{PV}{RT} = \frac{1.00 \text{ atm} \cdot 58.4 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 298 \text{ K}} = 2.38 \text{ mol}$
 $M = \frac{m}{n} = \frac{10.3 \text{ g}}{2.38 \text{ mol}} = 4.33 \text{ g/mol}$
 The gas is He .

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