

User's Guide

Huskylock 936



Huskylock 936 Threading Guide

Woodrow Wilson Chapman



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subsequent Δh values $2 \text{ CO}_2 \text{ g}$ $\text{H}_2\text{O g}$ $\text{C}_2\text{H}_2 \text{ g}$ $5 \text{ O}_2 \text{ g}$ $\text{C}_2\text{H}_2 \text{ g}$ $2 \text{ H}_2 \text{ g}$ $\text{C}_2\text{H}_6 \text{ g}$ $\Delta h 94.5 \text{ kJ}$ $\text{H}_2\text{O g}$ $\text{H}_2 \text{ g}$ $\frac{1}{2} \text{O}_2 \text{ g}$ $\Delta h 71.2 \text{ kJ}$ $\text{C}_2\text{H}_6 \text{ g}$ $7 \text{ O}_2 \text{ g}$ $2 \text{ CO}_2 \text{ g}$ $3 \text{ H}_2 \text{O g}$ $\Delta h 283 \text{ kJ}$ $\text{C}_2\text{H}_6 \text{ g}$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta h = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta h_{\text{products}}$

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exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions

1 $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$ ΔH_1

2 $\text{NaOH (s)} + \text{H}^+ \text{(aq)} \rightarrow \text{Na}^+ \text{(aq)} + \text{H}_2\text{O (l)}$ ΔH_2

3