

Assessment

Chapter Test B

Atomic Physics

MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- What is the frequency of a photon with an energy of $1.99 \times 10^{-19} \text{ J}$?
($h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$)
a. $1.00 \times 10^{14} \text{ Hz}$
b. $2.00 \times 10^{14} \text{ Hz}$
c. $3.00 \times 10^{14} \text{ Hz}$
d. $4.00 \times 10^{14} \text{ Hz}$
- Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?
a. 1 : 1
b. 2 : 1
c. 3 : 1
d. 4 : 1
- A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?
a. 4.6 eV
b. 4.8 eV
c. 1.4 eV
d. 2.4 eV
- What causes the bright lines in the emission spectrum of an element to occur?
a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.
- What causes the dark lines in the absorption spectrum of an element to occur?
a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

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VM Jensen



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